

PHYSICAL: Kinetic Molecular Theory

Learning Objectives

- At standard atmospheric conditions, all gases have approximately the same kinetic energy, but are made of different molecules and are made up of different numbers of particles. (The kinetic molecular theory is based on the assumption that all particles have the same average kinetic energy.)
- At standard atmospheric conditions, all gases have approximately the same average kinetic energy.

Model 1: Representation of Gases in Different Phases

The following figure illustrates the representation of particles in three phases: solid, liquid, and gas.



Figure 1: Representation of particles in different phases of matter.

At standard atmospheric conditions, all gases have approximately the same average kinetic energy, but are made of different molecules and are made up of different numbers of particles. (The kinetic molecular theory is based on the assumption that all particles have the same average kinetic energy.)

Key Concepts

1. All gases have approximately the same average kinetic energy at standard atmospheric conditions, but are made of different molecules and are made up of different numbers of particles. (The kinetic molecular theory is based on the assumption that all particles have the same average kinetic energy.)

Table 1: Representation of particles in different phases of matter.

	Solid	Liquid	Gaseous
Particle arrangement	Particles are in a fixed, regular arrangement. They are packed closely together. The particles are in a fixed position.	Particles are in a disordered, close-packed arrangement. They are packed closely together. The particles are in a disordered position.	Particles are in a highly disordered, spread-out arrangement. They are spread out. The particles are in a highly disordered position.
Particle movement	Particles are in a fixed position and can only vibrate. They are not free to move. The particles are in a fixed position and can only vibrate.	Particles are in a disordered position and can move. They are not free to move. The particles are in a disordered position and can move.	Particles are in a highly disordered position and can move. They are not free to move. The particles are in a highly disordered position and can move.

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David Kirk



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