

Name : \_\_\_\_\_ Date : \_\_\_\_\_

# Acids and Bases Strength

Write True (T) or False (F).

- i. Stronger acids have weaker conjugate bases and weaker acids have stronger conjugate bases. \_\_\_\_\_
- ii. Stronger bases have stronger conjugate acids and weaker bases have weaker conjugate acids. \_\_\_\_\_

Write equations that show how the following reactions occur in water.

- i.  $\text{F}^-$  acts as a base. \_\_\_\_\_
- ii.  $\text{HNO}_2$  acts as an acid. \_\_\_\_\_
- iii.  $\text{Fe}(\text{H}_2\text{O})_6^{3+}$  acts as an acid. \_\_\_\_\_
- iv.  $\text{HCO}_3^-$  acts as a base. \_\_\_\_\_
- v.  $\text{HCO}_3^-$  acts as an acid. \_\_\_\_\_
- vi.  $\text{Al}(\text{H}_2\text{O})_5(\text{OH})^{2+}$  acts as a base. \_\_\_\_\_

What volume of 0.45 M LiOH would be needed to neutralize 60.0 mL of 0.15 M HI?

Calculate the pH values of the following solutions of strong acids and bases.

- i. 0.001 M HCl
- ii. 0.76 M KOH
- iii.  $2.8 \times 10^{-4}$  M  $\text{Ba}(\text{OH})_2$

# Lecture 29 Acids And Bases I Worksheet

**DJ Losen**



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