

# Intro. to Chemical Reactions

Name \_\_\_\_\_

Date \_\_\_\_\_

Section \_\_\_\_\_

## Formula Review

1. Write the formulas for the following compounds:

- a. aluminum phosphate \_\_\_\_\_ b. sodium phosphate \_\_\_\_\_  
 c. copper (II) phosphate \_\_\_\_\_ d. sodium sulfate \_\_\_\_\_

2. Write the formulas for the following compounds:

- a.  $\text{CaCl}_2$  \_\_\_\_\_ b.  $\text{FeCl}_3$  \_\_\_\_\_  
 c.  $\text{FeCl}_2$  \_\_\_\_\_ d.  $\text{Cu}_2\text{S}$  \_\_\_\_\_

## Chemical Reactions: Virtual Energy Table

Write down observations for the reaction as it occurs:



There is no observable gas or color change when the two solutions are mixed.

Experiment 1 involves the reaction of hydrochloric acid with solid sodium hydroxide.

Write a balanced chemical equation for the reaction between sodium hydroxide and hydrochloric acid. Label your products.

## Chemical Table Data Collection

3. Write the following reaction under the correct label from above:



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**S Baum**



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