

10. Consider the following statements and choose the correct option from the given options [1]  
I. The time between the high tide and low tide, when the water level is falling, is called the ebb.  
II. The time between the low tide and high tide, when the tide is rising, is called the crest.

**Options:**

- (A) Only statement I is correct  
(B) Both the statements are correct  
(C) Only Statement II is correct  
(D) Both the statements I and II are incorrect
11. Which of the following is incorrectly matched? [1]

RIVER		ORIGIN
(A) Mahanadi	-	Near Sihawa in Raipur district of Chhattisgarh
(B) Godavari	-	Nasik district of Maharashtra
(C) Krishna	-	Near Mahabaleshwar in Sahyadri
(D) Narmada	-	Multa in the Betul district of Madhya Pradesh

12. The Chenab is an important tributary of [1]  
(A) The Ganga  
(B) The Kaveri  
(C) The Indus  
(D) The Narmada

13. **Assertion** With along coastline, large coastal areas have an equable climate. [1]

**Reason** The people of Mumbai and the Konkan coast have hardly any ideas of extremes of temperature and the seasonal rhythm of weather.

- (A) Only assertion is correct.  
(B) Only reason is correct.  
(C) Both assertion and reason are correct and reason is the correct explanation for assertion.  
(D) Both assertion and reason are correct but reason is not the correct explanation for assertion.
14. Which one among the following soil type is the result of intense leaching due to tropical rains? [1]  
(A) Laterite Soil  
(B) Alluvial Soil  
(C) Red Soil  
(D) Black Soil

**Read the following text and answer question no 15 to 17.**

In 1972, a comprehensive Wildlife Act was enacted, which provides the main legal framework for conservation and protection of wildlife in India. The two main objectives of the Act are; to provide protection to the endangered species listed in the schedule of the Act and to provide legal support to the conservation areas of the country classified as National parks, sanctuaries and closed areas. Thus Act has been comprehensively amended in 1991, making punishments more stringent and has also made provisions for the protection of specified plant species and conservation of endangered species of wild animals. There are 101 National parks and 553 wildlife sanctuaries in the country.

# Geography Grade 11 June Exam Theory

**Robert R. Redfield**



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web k sp of calcium hydroxide 1 author j m mccormick last update october 13 2013 introduction an equilibrium constant  $K$  is related to  $\Delta G$  for a process through eqn 1 and as long as  $\Delta H$  and  $\Delta S$  are independent of temperature eqn 1 can be combined with eqn 2 to determine  $\Delta H$  and  $\Delta S$  from the temperature dependence of  $K$

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web your primary objective in this experiment is to test a saturated solution of calcium hydroxide and use your observations and measurements to calculate the ksp of the compound you will do this by titrating the prepared  $\text{Ca(OH)}_2$  solution with a standard hydrochloric acid solution

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web in today's experiment you will determine the solubility product ksp of calcium hydroxide  $\text{Ca(OH)}_2$  by measuring the concentration of  $\text{Ca(OH)}_2$  in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of  $\text{Ca(OH)}_2$  is represented in equation form as shown below  $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$  the solubility product expression describes in mathematical terms the equilibrium

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web ksp for calcium hydroxide calcium hydroxide is a soft white caustic powder used in making mortar cements calcium salts paints and petrochemicals it is also used in saltwater aquaria to make up kalkwasser limewater solutions for reef tanks and is

used as a pH regulating agent notice that calcium hydroxide is divalent and thus has twice

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web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the solution  $\text{Ca(OH)}_2 \rightarrow \text{Ca}^{2+} + 2\text{OH}^-$   $5 \times 10^{-2} \text{ M}$  calcium ions  $5 \times 10^{-2} \text{ M}$   $0 \text{ M}$  hydroxide ions calculate calcium hydroxide ksp using the concentrations of calcium ions and hydroxide ions

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web madison bowe elana keller april 15 2014 introduction the objective of this lab was to measure the ksp of a slightly soluble basic salt and how solubility changes with temperature ksp is another form of equilibrium expression that deals with solubility the reaction that will take place is  $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$

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web the chemical equation used was  $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$  the solubility product was determined to be ksp  $\text{Ca}^{2+} \text{OH}^-$  from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

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web jul 12 2023 because the concentration of a pure solid such as  $\text{Ca}_3(\text{PO}_4)_2$  is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore  $\text{Ca}_3(\text{PO}_4)_2 \rightleftharpoons 3\text{Ca}^{2+} + 2\text{PO}_4^{3-}$   $K_{\text{sp}} = [\text{Ca}^{2+}]^3 [\text{PO}_4^{3-}]^2$

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web lab report 8 determining of calcium hydroxide the purpose of this was to titrate a saturated  $\text{Ca(OH)}_2$  solution with a standard HCl solution and find the ksp or the OH<sup>-</sup> concentration was found and then the concentration was found and u for this experiment a titration of calcium hydroxide with hydrochloric acid

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web question determining the ksp of calciumhydroxide  $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$  for this experiment the titration of calciumhydroxide with hydrochloric acid was investigated determining the ksp of calcium hydroxide  $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$  for this experiment the titration of calcium

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web question chem 182 experiment 8 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water solution of  $\text{Ca(OH)}_2$  is represented in equation form lution of  $\text{Ca(OH)}_2$  ionic solid that is sparingly soluble in water a saturated aqueous as shown below

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web therefore it was determined through error calculations and ease of use that the most reliable valid method to use to get the closest ksp value to the theoretical ksp for  $\text{Ca(OH)}_2$  was the determination through ph method 2 introduction calcium hydroxide is a strong base soluble in water

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