

## Chemical Kinetics

- 1) Molecular hydrogen can be made from methane gas by the reaction below. How is the rate of disappearance of  $\text{CH}_4$  related to the rate of appearance of  $\text{H}_2$ ?



- A) \_\_\_\_\_  
 B) \_\_\_\_\_  
 C) \_\_\_\_\_  
 D) none of these

Answer: B

- 2) For the reaction shown below, what is the relationship between the rate of formation of  $\text{Br}_2$  and the rate of reaction of  $\text{Br}^-$  over the same time period?



- A) rate of formation of  $\text{Br}_2 = 1/2 \times$  rate of reaction of  $\text{Br}^-$   
 B) rate of formation of  $\text{Br}_2 = 1/4 \times$  rate of reaction of  $\text{Br}^-$   
 C) rate of formation of  $\text{Br}_2 = 2 \times$  rate of reaction of  $\text{Br}^-$   
 D) rate of formation of  $\text{Br}_2 = 2 \times$  rate of reaction of  $\text{Br}^-$

Answer: A

- 3) The reaction that occurs in a Breathalyzer, a device used to determine the alcohol level in a person's bloodstream, is given below. If the rate of appearance of  $\text{Cr}_2(\text{SO}_4)_3$  is 1.34 mol/min at a particular moment, what is the rate of disappearance of  $\text{C}_2\text{H}_5\text{O}$  at that moment?



- A) 0.413 mol/min  
 B) 0.826 mol/min  
 C) 1.66 mol/min  
 D) 3.72 mol/min

Answer: C

# General Equilibrium Ap Chemistry Multiple Choice

**AN Whitehead**



**General Equilibrium Ap Chemistry Multiple Choice:**

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