

★ Raoult's law for a volatile liquid

$$P_1 = P_1^\circ x_1, \quad P_2 = P_2^\circ x_2$$

$$\text{Total pressure } P = P_1 + P_2 = P_1^\circ x_1 + P_2^\circ x_2$$

P_1 - partial vapour pressure of 1
 P_2 - partial vapour pressure of 2
 P_1° - vapour pressure of pure 1
 P_2° - vapour pressure of pure 2

★ Relative lowering of vapour pressure

$$\frac{P_1^\circ - P}{P_1^\circ} = \frac{W_2}{M_2} \times \frac{M_1}{W_1}$$

★ Elevation of boiling point

$$\Delta T_b = K_b m$$

K_b - ebullioscopic constant

$$\Delta T_b = \frac{K_b \times W_2 \times 1000}{M_2 \times W_1}$$

★ Depression of freezing point

$$\Delta T_f = K_f m$$

K_f - cryoscopic constant

$$\Delta T_f = \frac{K_f \times W_2 \times 1000}{M_2 \times W_1}$$

★ Osmotic pressure (π)

$$\pi = CRT$$

$$\pi = \frac{n_2}{V} RT$$

$$\pi = \frac{W_2 \times R \times T}{M_2 \times V}$$

W_2 - Mass of Solvent
 W_1 - Mass of Solute
 M_1 - Molar mass of solvent
 M_2 - Molar mass of solute
 C - Molarity
 T - Temperature

★ Van't Hoff factor (i)

$$i = \frac{\text{Normal molar mass}}{\text{Abnormal molar mass}}$$

$$i = \frac{\text{Observed colligative property}}{\text{Calculated colligative property}}$$

$$i = \frac{\text{Total no of moles of particles after association / dissociation}}{\text{no of moles of particles before association / dissociation}}$$

★ Degree of Dissociation

$$\alpha = \frac{i-1}{n-1}$$

n - number of ions

★ Modified form of Colligative properties

$$1) \frac{P_1^\circ - P_1}{P_1^\circ} = i \frac{n_2}{n_1}$$

$$2) \Delta T_b = i K_b m$$

$$3) \Delta T_f = i K_f m$$

$$4) \pi = i n_2 RT/V$$

Formula Sheet C12

JL Elias



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