Quiz 4 - Potential Energy Diagrams

- 1. The addition of a catalyst to a reaction provides an alternative mechanism with
 - A. Lower activation energy and lower reaction rate
 - B. Lower activation energy and higher reaction rate
 - C. Higher activation energy and lower reaction rate
 - D. Higher activation energy and higher reaction rate
- 2. Consider the following Reaction

$$\frac{1}{2} N_{2(g)} + \frac{1}{2} O_{2(g)} - NO_{(g)} = \frac{\Delta H}{2} = +90 \text{ kJ/mol NO}$$

The correct equation including the heat term is

- A. $N_{2(g)} + O_{2(g)} + 90 \text{ kJ } -----> 2NO_{(g)}$
- B. $N_{2(g)} + O_{2(g)} + 180 \text{ kJ} 2NO_{(g)}$
- C. $N_{2(g)} + O_{2(g)} - > 2NO_{(g)} + 90kJ$
- D. $N_{2(g)} + O_{2(g)} \cdots > 2NO_{(g)} + 180kJ$
- A forward reaction has an activation energy of 50 kJ and a ΔH of -100 kJ. The PE diagram, which describes this reaction, is

A.,

4. Consider the following potential energy diagram

The Activation energy for the forward reaction is

- A. 25 kJ
- B. 50 kJ
- C. 75 kJ
- D. 125 kJ
- 5. Consider the following reaction

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