



10. Which of the following are endothermic processes?

- A. ice melting
- B. water condensing
- C. freezing
- D. sublimation

11. Heat absorbed by a system at 25°C, positive temperature, is:

12. Which of the following is a statement of the first law of thermodynamics?

- A. the negative change in enthalpy corresponds to an exothermic process
- B. energy lost by the system must be gained by the surroundings, neither of which are created
- C. temperature is a measure of average kinetic energy
- D. enthalpy is a state function

13. The specific enthalpy of a closed system,  $H$ , is negative. Which of the following is true?

- A. Enthalpy is a state function, and the surroundings are not defined
- B. Enthalpy and the surroundings are not defined
- C. Enthalpy and the surroundings are defined
- D. Enthalpy and the surroundings are not defined
- E. None of these is a true statement

14. The reaction:



is \_\_\_\_\_, and therefore heat is \_\_\_\_\_ from the system.

- A. endothermic, released
- B. endothermic, absorbed
- C. exothermic, released
- D. exothermic, absorbed
- E. Thermodynamic values released and absorbed

15. Heat evolved in a reaction is observed when the reaction is carried out in a closed system at 25°C. The reaction is:



- A. 890.3 kJ
- B. 890.3 kJ
- C. 890.3 kJ
- D. 890.3 kJ
- E. 890.3 kJ



# Mcps Honors Chemistry Exam Guide

**R. R. Bowker LLC**



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