



# Leica 8x32 Ba Manual

**J Spring**



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web answer a 1 explanation the entire reaction was multiplied by 2 therefore  $\Delta h$  must also be halved 3 answer c explanation option c is incorrect because 566 kJ/mol is the  $\Delta h$  per mole of O<sub>2</sub> and not for 2 moles even though options b and d are expressed in grams they are still correct because 56

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the system to the surroundings  $2\text{H}_2\text{O}(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{l})$  energy endothermic the transfer of heat from the surroundings to the system energy  $2\text{H}_2\text{O}(\text{s}) \rightarrow 2\text{H}_2\text{O}(\text{g})$

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web  $q = 140\text{ g} \times 4.184\text{ J}/\text{K} \times 16.0^\circ\text{C}$   $q = 9372\text{ J}$   $35.5\text{ g} \times 1\text{ mole KCN} \times 0.5453\text{ moles}$   $65.1\text{ g} \times 0.5453\text{ mole moles}$  5 when a 15 g sample of  $\text{SrCl}_2$  dissolves in 80 g of water in a calorimeter the temperature drops from  $42^\circ\text{C}$  to  $15^\circ\text{C}$  calculate the  $\Delta h$  for the process

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