

- molar mass is for a compound (**g/mol**), e.g. $\text{CO}_2 = 44.01 \text{ g/mol}$
- atomic mass is for an element (**g/mol**), e.g. $\text{Al} = 26.98 \text{ g/mol}$

Scenario 1: calculating moles from grams

$$\text{grams of MX} \times \frac{1 \text{ mol of MX}}{\text{grams of MX}} \quad (\text{divide by the molar mass}); \text{ the molar mass isn't affected by the amount of moles in the chemical equation}$$

$$\text{grams of X} \times \frac{1 \text{ mol of X}}{\text{grams of X}} \quad (\text{divide by the atomic mass}); \text{ the atomic mass isn't affected by the amount of moles in the chemical equation}$$

$$88.02 \text{ g of CO}_2 \times \frac{1 \text{ mol CO}_2}{44.01 \text{ g of CO}_2} = 2 \text{ moles of CO}_2$$

Scenario 2: calculating grams from moles

$$\text{moles of MX} \times \frac{\text{grams of MX}}{1 \text{ mol of MX}} \quad (\text{multiply by the molar mass}); \text{ the molar mass isn't affected by the amount of moles in the chemical equation}$$

$$\text{moles of X} \times \frac{\text{grams of X}}{1 \text{ mol of X}} \quad (\text{multiply by the atomic mass}); \text{ the atomic mass isn't affected by the amount of moles in the chemical equation}$$

$$196.97 \text{ moles of Au} \times \frac{196.97 \text{ g of Au}}{1 \text{ mol of Au}} = 38797 \text{ g of Au}$$

Learning Guide To Stoichiometry

S Ashworth



Learning Guide To Stoichiometry:

Unveiling the Magic of Words: A Review of "**Learning Guide To Stoichiometry**"

In a global defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is really awe-inspiring. Enter the realm of "**Learning Guide To Stoichiometry**," a mesmerizing literary masterpiece penned by a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve into the book's central themes, examine its distinctive writing style, and assess its profound effect on the souls of its readers.

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