

Balancing Nuclear Equations

1. Write balanced equations for the following nuclear reactions:

- (A) The isotope C-11 decays by electron capture.
- (B) Curium-246 reacts with carbon-12 to produce nobelium-254 and four neutrons.
- (C) Californium-250 reacts with boron-10 to produce lawrencium-258 and two neutrons.
- (D) Fluorine-18 decays to oxygen-18 by positron emission.
- (E) Sodium-24 decays by beta emission.
- (F) Krypton-76 absorbs a beta particle to form bromine-76.

2. Fill in the blanks to complete the following nuclear reactions.

- (A) ${}^1_1\text{H}_1 + {}^7_3\text{Li}_3 \rightarrow$ _____
- (B) ${}^9_4\text{Be}_4 + {}^1_0\text{n}_0 \rightarrow {}^4_2\text{He}_2 +$ _____ $+ {}^4_2\text{He}_2$
- (C) ${}^{27}_{13}\text{Al}_{13} + {}^4_2\text{He}_2 \rightarrow$ _____ $+ {}^{30}_{15}\text{P}_{15}$
- (D) ${}^{63}_{29}\text{Cu}_{29} + {}^2_1\text{H}_1 \rightarrow 2 {}^1_0\text{n}_0 +$ _____
- (E) ${}^9_4\text{Be}_4 + {}^4_2\text{He}_2 \rightarrow$ _____ $+ {}^{12}_6\text{C}_6$
- (F) ${}^{239}_{93}\text{Np}_{93} \rightarrow {}^0_{-1}\text{e}_{-1} +$ _____
- (G) ${}^4_2\text{He}_2 + {}^{238}_{92}\text{U}_{92} \rightarrow 2 {}^1_0\text{n}_0 +$ _____

Key Terms Nuclear Changes Worksheet

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