

## Intro to Chemical Reactions:

### Test Review Answer Key



1. What do the symbols (s), (l), (g), and (aq) represent? (s) = solid, (l) = liquid, (g) = gas, and (aq) = aqueous
2. Why must chemical equations be balanced? Chemical equations must be balanced to satisfy the Law of Conservation of Mass. The number of atoms of the reactant side must equal the number of atoms on the product side of a chemical equation.
3. How can we be certain that a chemical reaction has taken place? Color change, temperature change, formation of a precipitate, production of gas (bubbles), and light emission.
4. Balance the following chemical equations:
  - a.  $1 \text{ AsCl}_3 + 3 \text{ H}_2\text{O} \rightarrow 3 \text{ HCl} + 1 \text{ As(OH)}_3$
  - b.  $4 \text{ Na} + 1 \text{ O}_2 \rightarrow 2 \text{ Na}_2\text{O}$
  - c.  $1 \text{ Cu} + 2 \text{ H}_2\text{SO}_4 \rightarrow 1 \text{ CuSO}_4 + 1 \text{ SO}_2 + 2 \text{ H}_2\text{O}$
  - d.  $2 \text{ Sb} + 3 \text{ H}_2\text{O} \rightarrow 1 \text{ Sb}_2\text{O}_3 + 3 \text{ H}_2$
  - e.  $1 \text{ La(NO}_3)_3 + 3 \text{ NaOH} \rightarrow 1 \text{ La(OH)}_3 + 3 \text{ NaNO}_3$
  - f.  $1 \text{ Ca(AlO}_2)_2 + 8 \text{ HCl} \rightarrow 2 \text{ AlCl}_3 + 1 \text{ CaCl}_2 + 4 \text{ H}_2\text{O}$
  - g.  $1 \text{ C}_4\text{H}_8 + 6 \text{ O}_2 \rightarrow 4 \text{ CO}_2 + 4 \text{ H}_2\text{O}$
  - h.  $2 \text{ C}_3\text{H}_8\text{O}_3 + 7 \text{ O}_2 \rightarrow 6 \text{ CO}_2 + 8 \text{ H}_2\text{O}$
5. Identify the following types of chemical reactions:
  - a.  $2 \text{ H}_2 + \text{O}_2 \rightarrow 2 \text{ H}_2\text{O}$  **Synthesis or Combustion**
  - b.  $2 \text{ H}_2\text{O} \rightarrow 2 \text{ H}_2 + \text{O}_2$  **Decomposition**
  - c.  $\text{Zn} + \text{H}_2\text{SO}_4 \rightarrow \text{ZnSO}_4 + \text{H}_2$  **Single Replacement**
  - d.  $2 \text{ CO} + \text{O}_2 \rightarrow 2 \text{ CO}_2$  **Synthesis or Combustion**
  - e.  $2 \text{ HgO} \rightarrow 2 \text{ Hg} + \text{O}_2$  **Decomposition**
  - f.  $2 \text{ KBr} + \text{Cl}_2 \rightarrow 2 \text{ KCl} + \text{Br}_2$  **Single Replacement**
  - g.  $\text{CaO} + \text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2$  **Synthesis**
  - h.  $\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$  **Double Replacement**
  - i.  $2 \text{ H}_2\text{O}_2 \rightarrow 2 \text{ H}_2\text{O} + \text{O}_2$  **Decomposition**
  - j.  $\text{Ca(OH)}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{CaSO}_4 + 2 \text{ H}_2\text{O}$  **Double Replacement**

# Intro To Chemical Reactions Test

**H Kauffman**



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