

Intro to Acids and Bases

Name: _____
 Class: _____
 Date: _____

Define each term: Definitions of acids and bases:

Arrhenius-definitions:

- acid substances that release dissolved cations hydrogen (H^+) (acids H^+ cations bonded to water as the hydronium ion, H_3O^+ , read H- three O H, C-2) and conjugated anions/bases
- base substances which release hydroxide ions: OH^- (base O H, C-2)

Brønsted-Lowry definitions:

- acid substances that donate a proton: H^+ , also accepted
- base substances which accept a proton: H^+ , also donated

Table 1: Representative reactions for H^+ (also known as the Lewis acid) and bases

Reaction	K_c
1. $HCl(aq) + H_2O(l) \rightleftharpoons H_3O^+(aq) + Cl^-(aq)$	1.0×10^6
2. $HNO_3 + H_2O \rightleftharpoons H_3O^+ + NO_3^-$ (strong acid)	4.0×10^9
3. $HNO_2(aq) + H_2O \rightleftharpoons H_3O^+ + NO_2^-$ (weak acid)	4.5×10^{-4}
4. $H_2SO_4 + H_2O \rightleftharpoons HSO_4^- + H_3O^+$	1.0×10^3
5. $H_2S + H_2O \rightleftharpoons HS^- + H_3O^+$	1.0×10^{-7}

Critical Thinking Questions

1. List acid characteristics for Table 1 that are characteristic acids:

are H^+ and OH^- ions

2. List acid characteristics for Table 1 that are characteristic bases:

are H^+ ions

3. List acid characteristics in Table 1 that are characteristic strong acids:

are HCl , HNO_3 , H_2SO_4 , H_2SO_3 , $HClO_4$, HNO_2

4. List acid characteristics in Table 1 that are characteristic weak acids:

are HNO_2 , H_2S , H_2SO_3

5. Is the generalization that all acids are bases? Explain:

Yes, because H^+ and OH^- ions are always present

6. Do acid substances contain hydroxide ions? Do they produce different reactions? Explain:

Yes, because H^+ is H^+ and OH^- is OH^- and they are not the same

7. Is this reaction involving acid and base? What is a hydroxide ion? What is a hydronium ion? Explain:

Intro To Acids And Bases Worksheet Answers

Chemquest 48

CL Gary



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