

# Intro to Acids and Bases

Name: \_\_\_\_\_  
 Class: \_\_\_\_\_  
 Date: \_\_\_\_\_

## Define each term: Definitions of acids and bases:

### Arrhenius Definitions:

- acid substances that release dissolved hydrogen ions ( $H^+$ ) (acids  $H^+$  gases formed in gases as the hydrogens form  $H_2$ ,  $HCl$ , and  $H_2$  (acids  $H_2$ ,  $HCl$ ) and compound substances)
- base substances which release hydroxide ions ( $OH^-$ )

### Brønsted-Lowry Definitions:

- acid substances that donate a proton ( $H^+$ ) (acid  $H^+$  gas)
- base substances which accept a proton ( $H^+$ ) (base  $H^+$  gas)

Table 1: Representative reactions for  $H^+$  (acid) and  $OH^-$  (base) species in aqueous solutions.

Reaction	Eq.
1. $HCl(aq) + H_2O(l) \rightarrow H_3O^+(aq) + Cl^-(aq)$	$HCl(aq) + H_2O(l)$
2. $HNO_3(aq) + H_2O(l) \rightarrow H_3O^+(aq) + NO_3^-(aq)$	$HNO_3(aq) + H_2O(l)$
3. $H_2SO_4(aq) + H_2O(l) \rightarrow H_3O^+(aq) + HSO_4^-(aq)$	$H_2SO_4(aq) + H_2O(l)$
4. $H_2SO_4(aq) + H_2O(l) \rightarrow H_3O^+(aq) + HSO_4^-(aq)$	$H_2SO_4(aq) + H_2O(l)$
5. $HCl(aq) + H_2O(l) \rightarrow H_3O^+(aq) + Cl^-(aq)$	$HCl(aq) + H_2O(l)$

## Critical Thinking Questions

- Are all the substances in Table 1 that are representative acids?
- Are all the substances in Table 1 that are representative bases?

### Class work:

- Are all the substances in Table 1 that are Brønsted-Lowry acids?

### Individual questions for students:

- Are all the substances in Table 1 that are Brønsted-Lowry bases?

### Classroom questions:

- Are all the substances in Table 1 that are Brønsted-Lowry acids?

### Individual questions for students:

- Are all the substances in Table 1 that are Brønsted-Lowry bases?

### Class:

- Are all the substances in Table 1 that are Brønsted-Lowry acids?

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**Edward Jasper King**



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