

## Part D Questions and Problems

26. An increase in the oxidation number of an atom indicates oxidation. A decrease in the oxidation number indicates reduction.
27. N is reduced (+5 to +2);  
Br is oxidized (-1 to 0)  
Mn is reduced (+7 to +2);  
Cl is oxidized (-1 to 0)  
N is reduced (+5 to +2);  
Sb is oxidized (0 to +5)  
S is reduced (+6 to +4);  
C is oxidized (0 to +4)

## Section Review 20.3

### Part A Completion

- |                     |          |
|---------------------|----------|
| 1. oxidation number | 5. two   |
| 2. half-reaction    | 6. added |
| 3. balanced         | 7. ionic |
| 4. ionic            |          |

### Part B True-False

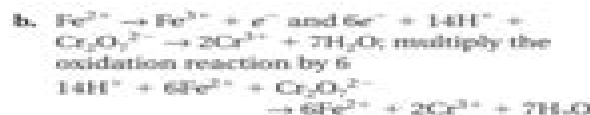
- |       |        |        |
|-------|--------|--------|
| 8. AT | 10. AT | 12. NT |
| 9. AT | 11. NT | 13. NT |

### Part C Matching

- |       |       |       |
|-------|-------|-------|
| 14. c | 17. b | 20. g |
| 15. a | 18. e |       |
| 16. f | 19. d |       |

## Part D Questions and Problems

21. a.  $\text{HNO}_3 \rightarrow \text{NO}$ ; N changes +5 to +2, a gain of  $3e^-$ ; multiply by 2  
 $2\text{HI} \rightarrow \text{I}_2$ ; I changes from -1 to 0, a loss of  $2e^-$  for  $\text{I}_2$ ; multiply by 3  
 $2\text{HNO}_3 + 6\text{HI} \rightarrow 2\text{NO} + 3\text{I}_2 + 4\text{H}_2\text{O}$
- b.  $\text{HNO}_3 \rightarrow \text{NO}_2$ ; N changes +5 to +4, a gain of  $1e^-$ ; multiply by 10  
 $\text{I}_2 \rightarrow 2\text{HIO}_3$ ; I changes from 0 to +5, a loss of  $10e^-$  for  $\text{I}_2$   
 $10\text{HNO}_3 + \text{I}_2 \rightarrow 2\text{HIO}_3 + 10\text{NO}_2 + 4\text{H}_2\text{O}$
22. a.  $\text{S}^{2-} \rightarrow \text{S} + 2e^-$  and  
 $3e^- + 4\text{H}^+ + \text{NO}_3^- \rightarrow \text{NO} + 2\text{H}_2\text{O}$   
Multiply the oxidation reaction by 3 and the reduction reaction by 2.  
 $3\text{H}_2\text{S} + 2\text{HNO}_3 \rightarrow 3\text{S} + 2\text{NO} + 4\text{H}_2\text{O}$



## Practice Problems

### Section 20.1

- |                        |                           |
|------------------------|---------------------------|
| 1. Sr:                 | oxidized (reducing agent) |
| $\text{O}_2$ :         | reduced (oxidizing agent) |
| 2. Li:                 | oxidized (reducing agent) |
| S:                     | reduced (oxidizing agent) |
| 3. Cs:                 | oxidized (reducing agent) |
| $\text{Br}_2$ :        | reduced (oxidizing agent) |
| 4. Mg:                 | oxidized (reducing agent) |
| $\text{N}_2$ :         | reduced (oxidizing agent) |
| 5. Fe:                 | oxidized (reducing agent) |
| $\text{O}_2$ :         | reduced (oxidizing agent) |
| 6. $\text{Br}^-$ :     | oxidized (reducing agent) |
| $\text{Cl}_2$ :        | reduced (oxidizing agent) |
| 7. Se:                 | oxidized (reducing agent) |
| $\text{F}_2$ :         | reduced (oxidizing agent) |
| 8. Ca:                 | oxidized (reducing agent) |
| $\text{O}_2$ :         | reduced (oxidizing agent) |
| 9. Mg:                 | oxidized (reducing agent) |
| $\text{H}^+$ :         | reduced (oxidizing agent) |
| 10. Na:                | oxidized (reducing agent) |
| $\text{H}_2\text{O}$ : | reduced (oxidizing agent) |

### Section 20.2

- a. Sn is tin in an uncombined state. The oxidation number is 0.
- b. The ionic charge on potassium is  $1+$ , thus the oxidation number is +1.
- c. The ionic charge on sulfur is  $2-$ , thus the oxidation number is -2.
- d. The ionic charge on iron is  $3+$ , thus the oxidation number is +3.
- e. Se is selenium in an uncombined state. The oxidation number is 0.
- f. The ionic charge on magnesium is  $2+$ , thus the oxidation number is +2.
- g. The ionic charge on tin is  $4+$ , thus the oxidation number is +4.

# Interpreting Graphics Answers Section Review

**S Ben Porath**



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