

**Solution**  
**CHAPTER TEST**  
**Class 11 - Chemistry**

1. (a) Propene

**Explanation:** Elimination reaction takes place, a molecule of HBr gets eliminated to form propene.

2. (c)  $\text{HI} > \text{HBr} > \text{HCl}$

**Explanation:**  $\text{HI} > \text{HBr} > \text{HCl}$

3. (a) 1,3-dimethylbenzene



4. (d) Geometric isomers

**Explanation:** Isomers which have the same structural formulae but differ in the relative spatial arrangements of atoms or groups around the double bond are called geometric isomers. Thus geometrical isomers are a type of stereoisomers.

5. (d)  $\text{HgSO}_4$

**Explanation:** Alkynes are not readily hydrated due to their low reactivity towards electrophilic addition reactions. Further dilute  $\text{H}_2\text{SO}_4$  has no effect on carbon carbon triple bond. In the presence of  $\text{HgSO}_4$  as catalyst, hydration occurs readily.

6. (a) A mixture of low boiling hydrocarbons.

**Explanation:** Pyrolysis is the decomposition of a compound by the application of heat. It involves the breaking of C-C and C-H bonds resulting in the formation of lower boiling alkanes.

7. (a) 3

**Explanation:** The position isomers of  $\text{C}_5\text{H}_8$  are n-pentyne, 3-methyl butyne, pent-2-yne.

8. (a)  $\text{C}_8\text{H}_{16}$

**Explanation:** Alkenes are unsaturated hydrocarbons with at least one double bond between the carbon atoms. Their general formula is  $\text{C}_n\text{H}_{2n}$ .

9. (b) Sodium acetate

**Explanation:**

This is an example of Kolbe's electrolysis method. The reaction is:



The step-wise redox reactions occurring in the electrolytic cell are depicted as under



At Anode: (Oxidation)



At Cathode: (Reduction)



# Hydrocarbons Answer Key

**J Spring**



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