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oxidation $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^{-}$ reduction $\text{Fe}^{2+} + 2\text{e}^{-} \rightarrow \text{Fe}$ $24 \text{ 3 Ag} + \text{Ni} \rightarrow \text{Ni} + 3 \text{ Ag}$ oxidation reduction $25 \text{ Cu} + \text{Pb} \rightarrow \text{Pb} + \text{Cu}$ oxidation reduction 26 O_2

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 $\text{b} \text{ 5CO} + \text{I}_2 \rightarrow \text{5CO}_2 + \text{I}_2$
 $\text{c} \text{ O}_2 + \text{C} \rightarrow \text{CO}_2$

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web jun 28 2010 answers i determine what is oxidized and what in each reaction identify the oxidizing agent and the reducing agent also $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$ $\text{Sr} \rightarrow \text{Sr}^{2+}$ oxidized reducing agent $\text{O} \rightarrow \text{O}^{2-}$ reduced $\text{Ox ag} \text{ 2Li} + \text{s} \rightarrow \text{Li}_2\text{s}$ $\text{Li} \rightarrow \text{Li}^{+}$ oxidized red ag

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