

**PEMBAHASAN SOAL PILIHAN GANDA SELEKSI TIM OSN KIMIA TK PROVINSI 2013**  
**Oleh Urip Kalteng (<http://urip.wordpress.com>) pada 6 Juni 2013**

1. Padatan berikut ini, manakah atom-atom atau molekul-molekulnya yang berinteraksi hanya oleh gaya van der Waals?
- $\text{CO}_2$
  - $\text{SiO}_2$
  - Cs
  - $\text{MgO}$
  - $\text{CH}_3\text{CH}_2\text{OH}$

**Pembahasan:**

Zat yang bersifat non polar hanya memiliki interaksi gaya van der Waals antarmolekulnya. Di antara spesies pada alternatif jawaban yang bersifat non polar hanyalah  $\text{CO}_2$ .

2. Dari pernyataan berikut ini, manakah yang menggambarkan fenomena adanya ikatan hidrogen antar-molekul?
- Titik leleh dari senyawa hidroksida golongan 1 meningkat dengan bertambahnya massa molekul relatif ( $M_r$ )
  - Titik didih alkana meningkat dengan bertambahnya massa molekul relatif
  - $\text{CH}_3\text{OCH}_3$  ( $M_r$ : 46) mempunyai titik didih lebih tinggi dari  $\text{CH}_3\text{CH}_2\text{CH}_3$  ( $M_r$ : 44)
  - Hidrogen klorida membentuk suatu larutan asam bila dilarutkan dalam air
  - Pada  $0^\circ\text{C}$ , es mempunyai densitas lebih rendah dari air

**Pembahasan:**

Fenomena adanya ikatan hidrogen antarmolekul yang paling tepat ditunjukkan pada fenomena  $\text{H}_2\text{O}$  yang ketika dalam bentuk es pada suhu  $0^\circ\text{C}$  memiliki kerapatan (densitas) yang lebih rendah dibandingkan densitas air.

3. Jari-jari dan muatan masing-masing dari enam ion ditunjukkan dalam tabel ini:

| Ion            | J    | L    | M <sup>+</sup> | N    | Y <sup>-</sup> | Z <sup>2-</sup> |
|----------------|------|------|----------------|------|----------------|-----------------|
| Jari-jari (nm) | 0,14 | 0,15 | 0,15           | 0,14 | 0,15           | 0,15            |

Senyawa padatan ionik JX, LY, dan MZ masing-masing mempunyai jenis kisi yang sama. Manakah urutan yang benar dari penembapatan energi kisinya mulai dari yang tertinggi hingga terendah?

- $\text{JX} > \text{LY} > \text{MZ}$
- $\text{JX} > \text{MZ} > \text{LY}$
- $\text{LY} > \text{MZ} > \text{JX}$
- $\text{MZ} > \text{JX} > \text{LY}$
- $\text{MZ} > \text{LY} > \text{JX}$

**Pembahasan:**

Energi kisi secara sederhana dapat diperoleh dari energi potensial elektrostatis antara kation dengan anion melalui persamaan:

$$E = k \frac{|q_1 q_2|}{r} \quad \text{Nilai } E \text{ akan semakin besar jika nilai } \frac{|q_1 q_2|}{r} \text{ makin besar.}$$

|                         | JX                   | LY                   | MZ                   |
|-------------------------|----------------------|----------------------|----------------------|
| Jarak kation-anion (nm) | $0,14 + 0,14 = 0,28$ | $0,15 + 0,15 = 0,30$ | $0,15 + 0,15 = 0,30$ |
| Muatan kation-anion     | +1 dan -1            | +1 dan -1            | +2 dan -2            |
| $\frac{ q_1 q_2 }{r}$   | 3,57                 | 2,78                 | 13,3                 |

# Hasil Osn Kalteng

**Lauren Gardner**



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