

## Chapter 2: Atoms, Molecules and Ions

### Dalton's Atomic Theory (1805)

1. All matter consists of **atoms**; tiny indivisible particles of an element that cannot be created or destroyed.
2. Atoms of one element **cannot** be subdivided, created, or destroyed during chemical reactions. (wrong)
3. Atoms of an element are **identical** in mass and other properties and are different from the atoms of any other element. (wrong)
4. Compounds result from the chemical combination of a **specific ratio** of atoms of different elements.

### Atomic Structure: Electrons

- J.J. Thomson (1897) Cathode Ray Tube Experiment
- Observations:
  - Ray bends in magnetic field
  - Ray bends toward positive plate in electric field
  - Ray is identical for any cathode
- Conclusions:
  - Ray consists of charged particles
  - Ray consists of negative particles
  - These particles are found in ALL matter

**Cathode Ray Tube experiment using magnets led to the discovery of electrons**

### Thomson's Model of the Atom

- Plum-pudding model:
  - E- distributed throughout diffuse, positively charged "squishy" sphere



### Atomic Structure: The Nucleus

- Rutherford's Experiment
  - Bombard a thin gold foil with positively charged particles to test Thomson's model of the atom
  - Observations:
    - Atom is made up of smaller subatomic particles
      - Protons
      - Neutrons
      - Electrons (mass of only ab 1/2000th of other particles)
  - Conclusions

# Genral Chemistry Study Guide

**Matthew Bowling**



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## **Table of Contents Genral Chemistry Study Guide**

1. Understanding the eBook Genral Chemistry Study Guide
  - The Rise of Digital Reading Genral Chemistry Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Genral Chemistry Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Genral Chemistry Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Genral Chemistry Study Guide
  - Personalized Recommendations
  - Genral Chemistry Study Guide User Reviews and Ratings
  - Genral Chemistry Study Guide and Bestseller Lists
5. Accessing Genral Chemistry Study Guide Free and Paid eBooks
  - Genral Chemistry Study Guide Public Domain eBooks
  - Genral Chemistry Study Guide eBook Subscription Services
  - Genral Chemistry Study Guide Budget-Friendly Options

6. Navigating Genral Chemistry Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Genral Chemistry Study Guide Compatibility with Devices
  - Genral Chemistry Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Genral Chemistry Study Guide
  - Highlighting and Note-Taking Genral Chemistry Study Guide
  - Interactive Elements Genral Chemistry Study Guide
8. Staying Engaged with Genral Chemistry Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Genral Chemistry Study Guide
9. Balancing eBooks and Physical Books Genral Chemistry Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Genral Chemistry Study Guide
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Genral Chemistry Study Guide
  - Setting Reading Goals Genral Chemistry Study Guide
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Genral Chemistry Study Guide
  - Fact-Checking eBook Content of Genral Chemistry Study Guide
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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