Chapter 2: Atoms, Molecules and Ions

Dalton's Atomic Theory (1805)

- All matter consists of atoms; tiny indivisible particles of an element that cannot be created or destroyed.
- 2. Atoms of one element cannot be subdivided, created, or destroyed during chemical reactions (ipmortion)
- 3. Atoms of an element are identical in mass and other properties and are different from the atoms of any other element. (wnonw)
- 4. Compounds result from the chemical combination of a specific ratio of atoms of different elements.

Atomic Structure: Flectrons

- J.J. Thomson (1897) Cathode Ray Tube Experiment
- Observations:
 - Ray bends in magnetic field
 - Ray bends toward positive plate in electric field
 - Ray is identical for any cathode
- Conclusions:
 - Ray consists of charged particles
 - Ray consists of negative particles
 - These particles are found in ALL matter

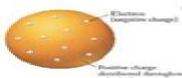
Cathode Ray Tube experiment using magnets led to the discovery of electrons

Thomson's Model of the Atom

- Plum-pudding model:
 - E- distributed throughout diffuse, positively charged "squishy" sphere

Atomic Structure: The Nucleus

- Rutherford's Experiment
 - Bombard a thin gold foil with positively charged particles to test Thomson's model of the atom.
 - Observations:
 - Atom is made up of smaller subatomic particles
 - Protons
 - Neutrons
 - Electrons (mass of only ab 1/2000th of other particles)
 - Conclusions



Genral Chemistry Study Guide

Matthew Bowling

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