

PROPERTIES OF MATTER STUDY GUIDE



ATOMS, ELEMENTS, COMPOUNDS + MIXTURES

Complete the vocabulary matching below by drawing a line to match the correct definition to the correct vocabulary word.

atom	→	d. A mixture in which the different parts can be seen because the substances are not evenly distributed.
element	→	b. The basic unit of matter that is made of protons, neutrons and electrons.
compound	→	c. A mixture in which the substances are evenly distributed and you cannot see the different parts.
homogeneous mixture	→	e. A pure substance that is made of 2 or more atoms that are chemically combined.
heterogeneous mixture	→	a. A pure substance that is made of only 1 type of atom and can be found on the periodic table.

Complete the chart below about the subatomic particles and their charges:

	PARTICLE #1	PARTICLE #2	PARTICLE #3
NAME	proton	neutron	electron
CHARGE	+ (positive)	0 (neutral)	- (negative)
LOCATION	part of nucleus	part of nucleus	outer "cloud-like" orbit

Complete the graphic organizer below about matter below. Need help? Use the word bank to assist you!

heterogeneous	element	mixture	compound
	pure substance	electron	homogeneous



pure substance is composed of



carbon (C)

hydrogen (H)

oxygen (O)

nitrogen (N)

fluorine (F)

chlorine (Cl)

iodine (I)

bromine (Br)

sulfur (S)

phosphorus (P)

silicon (Si)

germanium (Ge)

tin (Sn)

lead (Pb)

mixture

homogeneous

heterogeneous

pure substance

element

compound

mixture

pure substance

element

compound

mixture

pure substance

element

compound

Classify the following examples as elements, compounds or mixtures in the table below.

carbon (C)	salt	water (H ₂ O)
hydrogen (H)	salt water	sugar (sucrose)

ELEMENT	COMPOUND	MIXTURE
carbon (C) hydrogen (H)	water (H ₂ O) salt (NaCl) sugar (sucrose)	salt

Why is carbon considered an element and carbon dioxide considered a compound?

Carbon is an element as the periodic table, made of only carbon atoms. Carbon dioxide is carbon and oxygen chemically combined together.

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