

## Topic 10.1: Organic chemistry – Fundamentals of organic chemistry

Organic chemistry focuses on the chemistry of compounds containing carbon.

• **Understanding:** A homologous series is a series of compounds of the same family, with the same general formula, which differ from each other by a common structural unit.

- **Homologous series:** series of compound of the same family, with the same general formula, which differ by a common structural unit
  - Common structural unit: often  $\text{CH}_2$  (methylene) group
  - Same family: has the same functional group
  - Properties: homologous series has similar physical and chemical properties

• **Understanding:** Structural formulas can be represented in full and condensed format.

|             | Full structural formula                                                              | Condensed structural formula                                                                            | Skeletal structural formula                                                                                                           |
|-------------|--------------------------------------------------------------------------------------|---------------------------------------------------------------------------------------------------------|---------------------------------------------------------------------------------------------------------------------------------------|
| Example     |     | $\text{CH}_3\text{—CH}_2\text{—CH}_2\text{—CH}_3$<br><br>$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$ |                                                    |
| Description | Two-dimensional representation showing all atoms, their relative positions and bonds | Two-dimensional representation showing all atoms and their relative position                            | Two dimensional representation of a compound where each vertex and end of line represent a carbon atom, and functional group is added |

• **Understanding:** Structural isomers are compounds with the same molecular formula but different arrangements of atoms.

- **Structural isomers:** compounds with same molecular formula with different arrangement of atoms
  - Properties: structural isomers have different chemical and physical properties (melting point, boiling point etc.)

• **Understanding:** Functional groups are the reactive parts of molecules.

- **Functional groups:** reactive parts of a molecule that is responsible for the molecule's chemical property
  - Naming: the position of the function group is identified by the number of carbon to which it is attached

• **Understanding:** Saturated compounds contain single bonds only and unsaturated compounds contain double or triple bonds.

- **Saturated compounds:** organic compounds in which all the carbon-carbon bonds are single (e.g., *alkanes*)
- **Unsaturated compounds:** organic compound in which there is at least one carbon-carbon bond that is double or triple (e.g., *alkenes*)

• **Understanding:** Benzene is an aromatic, unsaturated hydrocarbon.

- **Aliphatic:** characterized by a straight, branched or cyclic structure
- **Aromatic:** characterized by presence of benzene ring
- **Benzene:** aromatic, unsaturated hydrocarbon in which six carbons bonds into a ring
  - **Sigma ( $\sigma$ ) bond in benzene:** benzene contains six  $\text{sp}^2$  hybridized carbon atoms bond to one other by sigma bonds
  - **Pi ( $\pi$ ) bond in benzene:** p-orbitals in carbon overlap to form a continuous  $\pi$  bond above and below the plane of carbons

• **Applications and skills:** Explanation of the trends in boiling points of members of a homologous series.

| Name    | Formula                   | Structural formula (condensed)                                       | Boiling point ( $^{\circ}\text{C}$ ) |
|---------|---------------------------|----------------------------------------------------------------------|--------------------------------------|
| Methane | $\text{CH}_4$             | $\text{CH}_4$                                                        | -161                                 |
| Ethane  | $\text{C}_2\text{H}_6$    | $\text{CH}_3\text{CH}_3$                                             | -89                                  |
| Propane | $\text{C}_3\text{H}_8$    | $\text{CH}_3\text{CH}_2\text{CH}_3$                                  | -42                                  |
| Butane  | $\text{C}_4\text{H}_{10}$ | $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$                       | -0.5                                 |
| Pentane | $\text{C}_5\text{H}_{12}$ | $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$            | 36                                   |
| Hexane  | $\text{C}_6\text{H}_{14}$ | $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$ | 69                                   |

- Trend in boiling points: boiling points increases as the carbon chain increases by a common structural unit
  - Trends arise from increasing intermolecular (dispersion forces) with increased chain lengths

• **Applications and skills:** Distinction between empirical, molecular and structural formulas.

- **Empirical formula:** simplest ratio of atoms (of an element) in a molecule
- **Molecular formula:** actual number of atoms (of each element) in a molecule
- **Structural formula:** representation of the type of atom and their position relative to one other

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