

# I've got the Power! The Voltaic Cell

## *Why?*

How exactly does a battery work? The tendency for electrons to flow from one substance to another is something that can be channeled and controlled. A voltaic (or electrochemical) cell provides a mechanism for the conversion of chemical energy into electrical energy. The cell potential, or voltage, and current that are produced by an electrochemical cell are determined by the chemicals that comprise the cell. In this activity, you will see that the amount of voltage produced by a cell depends upon the chemical composition.

## *Learning Objectives*

- Understand how the cell potential is related to composition of the cell.
- See how measurements of cell potentials are used to construct an activity series.
- Determine whether a redox reaction is spontaneous or non-spontaneous

## *Success Criteria*

- Demonstrate an understanding of the operation of a voltaic cell.

## *Prerequisites*

- Ability to write oxidation and reduction half-reactions
- Introduction to Voltaic Cell diagrams and terminology

## *Requirements*

- Activity Series reference table
- Computer and internet access for the following site:  
<http://www.blackgold.ab.ca/ict/Division4/Science/Div.%204/Voltaic%20Cells/Voltaic.htm>

## *Vocabulary*

- |                |             |
|----------------|-------------|
| • Reduction    | • Oxidation |
| • Voltaic cell | • Half-cell |
| • Anode        | • Cathode   |
| • Salt bridge  |             |

## *Information*

Cell potentials for spontaneous voltaic cells are positive. A voltmeter reading can be positive or negative depending on which wires are connected to which electrodes. If a voltmeter gives a negative reading for voltage, switching the connections between the wires and the electrodes will produce a positive voltmeter reading.

# I Ve Got The Power The Voltaic Cell Answers

**Arihant Experts**



## **I Ve Got The Power The Voltaic Cell Answers:**

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