

⇒ Heisenberg uncertainty principle :

$$\Delta x \cdot \Delta p \geq \frac{h}{4\pi}$$

⇒ Number of orbitals in a shell = n^2

⇒ Number of electrons in a shell = $2n^2$

⇒ Number of radial nodes = $(n - l - 1)$

Chemical Bonding

$$H = \frac{1}{2}[V + X - C + A]$$

(H = number of Hybrid orbitals)

$$\Rightarrow \text{Bond order} = \frac{1}{2}[N_b - N_a]$$

States of matter

⇒ Boyle's law : $P_1 V_1 = P_2 V_2$

⇒ Charles law : $\frac{V_1}{T_1} = \frac{V_2}{T_2}$

⇒ Ideal gas equation : $PV = nRT$

Dalton's law : $\Rightarrow P_{\text{total}} = P_1 + P_2 + P_3 + \dots$

$$\Rightarrow d_{\text{gas}} = \frac{PM}{RT}$$

$$\text{Graham's Law} \Rightarrow \frac{r_1}{r_2} = \sqrt{\frac{M_2}{M_1}}$$

⇒ Vander waal's equation : $\left(P + \frac{an^2}{V^2}\right)(V - nb) = nRT$

⇒ $C_p - C_v = nRT$

⇒ Kinetic gas equation : $PV = \frac{1}{3}Mu^2$

Chemical Thermodynamics

⇒ $\Delta U = q_p$, $\Delta H = q_p$

⇒ $\Delta H = \Delta U + \Delta n_g RT$

⇒ $C = \frac{q}{\Delta T}$ (C = Heat capacity)

⇒ $\Delta U = q + w$ (First law)

$$\Rightarrow \Delta H = \sum \Delta_f H^\circ_{\text{products}} - \sum \Delta_f H^\circ_{\text{reactants}}$$

⇒ $\Delta G = \Delta H - T\Delta S$

⇒ $\Delta G^\circ = -RT \ln K$

Equilibrium

$$\Rightarrow K_{eq} = \frac{K_c}{K_p}$$

⇒ $K_p = K_c(RT)^{\Delta n}$

⇒ $K_w = [H^+][OH^-]$

⇒ $pH = -\log [H^+]$

⇒ $pK_w = pH + pOH$

⇒ $K_a = C\alpha^2$

⇒ $K_w = K_a \times K_b$

⇒ $K_w = [A^-]_e [B^+]_e$

Solid State

$$\Rightarrow \text{Density} = \frac{Z \times M}{a^3 \times N_A} \text{ g/cm}^3$$

⇒ Packing fraction : Simple cubic = $\frac{\pi}{6}$, FCC = $\frac{\sqrt{2}\pi}{6}$, BCC = $\frac{\sqrt{3}\pi}{8}$

⇒ Bragg's equation $d = n\lambda \sin \theta$

Solution

⇒ Henry's law $X_{\text{gas}} = K_H P$

⇒ Raoult's law : v.p. of solution = $x_A P_A^\circ + x_B P_B^\circ$

⇒ Relative lowering of V.P. : $\frac{P_A^\circ - P_A}{P_A^\circ} = X_{\text{solvent}}$

⇒ $\Delta T_b = i K_b m$

⇒ $\Delta T_f = i K_f m$

⇒ $\pi = CRT$

⇒ $i = \frac{\text{no. of particle of solute in solution}}{\text{no. of particles of pure solute}}$

Electrochemistry

$$\Rightarrow E^\circ_{\text{cell}} = E^\circ_{\text{cathode}} - E^\circ_{\text{anode}}$$

$$\Rightarrow \Delta G^\circ = -nF E^\circ_{\text{cell}}$$

Hsc Chemistry Formula Sheet

Y Pai



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