#### GAS STOICHIOMETRY PROBLEMS WORKSHEET I

1. Given the unbalanced decomposition reaction of baking soda:

2NaHCO<sub>3 to</sub> + heat ===> Na<sub>3</sub>CO<sub>3 to</sub> + CO<sub>2 to</sub> + HOH to 42.0 grams of

baking soda? What volumes of carbon dioxide and water are produced at STP?

The catalytic decomposition of hydrogen peroxide is:

2H<sub>2</sub>O<sub>2 (set)</sub> ===> 2H<sub>2</sub>O<sub>30</sub> + O<sub>2 (et)</sub>
Balance the reaction. How many moles of water and oxygen are produced by the decomposition of 68.0 grams of hydrogen peroxide? How many molecules of water and oxygen are produced? How many grams of each product are formed?

If 156.0 grams of potassium metal reacts with excess water.potassium hydroxide and hydrogen gas are formed? What volume of hydrogen gas, in liters.

Be sure to balance the reaction first.  $2K_{10} + 2H_{2}O_{10} \rightarrow 2KOH_{100} + H_{200}$   $1.5\% J \times \sqrt{\frac{1 \text{ mod fit}}{32.00 \text{ g fc}}} \times \frac{1 \text{ mod fit}}{2 \text{ mod fit}} \times \sqrt{\frac{12.4\% \text{ g fc}}{1 \text{ mod fit}}} \times \frac{444.26 \text{ g fc}}{20.20 \text{ g fc}}$   $1.5\% J \times \sqrt{\frac{1 \text{ mod fit}}{32.00 \text{ g fc}}} \times \frac{2 \text{ mod fit}}{2 \text{ mod fit}} \times \frac{22.4\% \text{ g fc}}{2 \text{ mod fit}} \times \frac{27.4\% \text{ g fc}}{20.20 \text{ g fc}} \times \frac{27.4\% \text{ g fc}}{2 \text{ mod fit}} \times \frac{2$ 

4. Determine the number of moles of carbon dioxide gas, water, and sodium chloride formed by the reaction of 42.0 grams of sodium bicarbonate (baking soda) reacting with excess hydrochloric acid. The reaction is:

NaHCO<sub>3 (4)</sub> + HCI<sub>1000</sub> ===> CO<sub>3 (6)</sub> + HOH<sub>101</sub> + NaCI<sub>1000</sub>

423 NoHCO<sub>3 (4)</sub> = 1 rel CO<sub>3 (4)</sub> = 0.5 rel CO<sub>3</sub>

NaHCO<sub>3 (4)</sub> = 1 rel NoHCO<sub>3</sub> = 0.5 rel CO<sub>3</sub>

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