

12 • The Gas Laws**CHARLES'S LAW**

Charles' Law states the volume of a gas varies directly with the Kelvin temperature, assuming the pressure is constant. We use the following formulas:

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \quad \text{or} \quad V_1 \times T_2 = V_2 \times T_1$$

$$K = ^\circ C + 273$$

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Solve the following problems assuming a constant pressure. Assume all numbers are 3 significant figures.

1. A sample of nitrogen occupies a volume of 250 mL at $25^\circ C$. What volume will it occupy at $95^\circ C$?

$$\begin{matrix} V_1 & T_1 & T_2 \\ & 25^\circ C & 95^\circ C \\ & 298 K & 368 K \end{matrix}$$
2. Oxygen gas is at a temperature of $40^\circ C$ when it occupies a volume of 2.30 Liters. To what temperature should it be raised to occupy a volume of 6.50 Liters?

$$\begin{matrix} T_1 & V_1 & V_2 \\ 40^\circ C & 2.30 L & 6.50 L \\ 313 K & & \end{matrix}$$
3. Hydrogen gas was cooled from $150^\circ C$ to $50^\circ C$. Its new volume is 75.0 mL. What was its original volume?

$$\begin{matrix} T_1 & T_2 & V_2 \\ 150^\circ C & 50^\circ C & 75.0 mL \\ 423 K & 323 K & \end{matrix}$$
4. Chlorine gas occupies a volume of 25.0 mL at 300 K. What volume will it occupy at 600 K?

$$\begin{matrix} V_1 & T_1 & T_2 \\ 25.0 mL & 300 K & 600 K \end{matrix}$$
5. A sample of neon gas at $50^\circ C$ and a volume of 2.50 Liters is cooled to $25^\circ C$. What is the new volume?

$$\begin{matrix} T_1 & V_1 & T_2 \\ 50^\circ C & 2.50 L & 25^\circ C \\ 323 K & & 298 K \end{matrix}$$
6. Fluorine gas at 300 K occupies a volume of 500 mL. To what temperature should it be lowered to bring the volume to 300 mL?

$$\begin{matrix} T_1 & V_1 & V_2 \\ 300 K & 500 mL & 300 mL \end{matrix}$$
7. Helium occupies a volume of 3.80 Liters at $-45^\circ C$. What volume will it occupy at $45^\circ C$?

$$\begin{matrix} V_1 & T_1 & T_2 \\ 3.80 L & -45^\circ C & 45^\circ C \\ & 228 K & 318 K \end{matrix}$$
8. A sample of argon gas is cooled and its volume went from 380 mL to 250 mL. If its final temperature was $-55^\circ C$, what was its original temperature?

$$\begin{matrix} V_1 & V_2 & T_2 \\ 380 mL & 250 mL & -55^\circ C \\ & & 218 K \end{matrix}$$

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