

Gas Laws Problem Packet

Gas Laws Packet

Name _____ Date _____ Period _____

Instructions

Standard Conditions

Unit	Standard Pressure
atm	1.013
torr	760
mm	760
cm	760
in	30

22.4 L = 1 mol at STP

STP (Standard Temperature and Pressure) = 0°C and 1 atm

The gas laws:

Boyle's Law	Charles's Law
$P_1V_1 = P_2V_2$	$V_1/T_1 = V_2/T_2$
Combined Gas Law	Avogadro's Law
$P_1V_1/T_1 = P_2V_2/T_2$	$V_1/n_1 = V_2/n_2$
Avogadro's Law	Combined Gas Law
$V_1/n_1 = V_2/n_2$	$P_1V_1/T_1 = P_2V_2/T_2$
Combined Gas Law	Avogadro's Law
$P_1V_1/T_1 = P_2V_2/T_2$	$V_1/n_1 = V_2/n_2$
Avogadro's Law	Combined Gas Law
$V_1/n_1 = V_2/n_2$	$P_1V_1/T_1 = P_2V_2/T_2$

The constant R (used in the ideal gas law) depends on the unit of pressure:

- If kPa, $R = 8.314 \frac{\text{J}}{\text{mol} \cdot \text{K}}$
- If atm, $R = 0.0821 \frac{\text{atm} \cdot \text{L}}{\text{mol} \cdot \text{K}}$

Notes:

- Temperature must always be in Kelvin! $K = ^\circ\text{C} + 273$
- If you are given two different units of pressure, CONVERT (one).
- Create two columns before solving: initial and final. Write what you're given in the problem. Find the law that uses the variables you're given, and solve for the unknown. (for Ideal Gas Law, write what you are given, insert those in equation and solve for unknown)
- Avogadro's Law: DON'T FORGET TO USE MOLE (Amount) (always)
- Graham's Law: You will ALWAYS have the molecular mass for these problems. Double and triple check you have the rate of effusion and molecular mass in the correct spots before solving the equation (this is an INVERSE relationship). You MUST calculate the square root of double both sides of the equation before moving things.

Charles' Law $V_1/T_1 = V_2/T_2$

1. The temperature under my refrigerator is about 4°C. If I place it below the temperature of 4°C, and it is below 4°C, what will happen to the volume of the refrigerator?

2. A balloon is at 20°C.

Boyle's Law $P_1V_1 = P_2V_2$

1. A gas is at 1.0 atm and 2.0 L. If the pressure is increased to 2.0 atm, what will happen to the volume?

2. A gas is at 1.0 atm and 2.0 L. If the pressure is decreased to 0.5 atm, what will happen to the volume?

3. A gas is at 1.0 atm and 2.0 L. If the pressure is increased to 2.0 atm, what will happen to the volume?

4. A gas is at 1.0 atm and 2.0 L. If the pressure is decreased to 0.5 atm, what will happen to the volume?

5. A gas is at 1.0 atm and 2.0 L. If the pressure is increased to 2.0 atm, what will happen to the volume?

6. A gas is at 1.0 atm and 2.0 L. If the pressure is decreased to 0.5 atm, what will happen to the volume?

7. A gas is at 1.0 atm and 2.0 L. If the pressure is increased to 2.0 atm, what will happen to the volume?

8. A gas is at 1.0 atm and 2.0 L. If the pressure is decreased to 0.5 atm, what will happen to the volume?

9. A gas is at 1.0 atm and 2.0 L. If the pressure is increased to 2.0 atm, what will happen to the volume?

10. A gas is at 1.0 atm and 2.0 L. If the pressure is decreased to 0.5 atm, what will happen to the volume?

All Gas Laws Included!



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Arshad Iqbal



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Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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