

# Gas Laws Practice

Name: \_\_\_\_\_

Date: \_\_\_\_\_

Hour: \_\_\_\_\_

**IMPORTANT:** whenever you use temperature, it must be in degree Kelvin (K), so remember the equation:  $K = ^\circ C + 273$

1. a) convert  $39\text{ }^\circ\text{C}$  to K.                      b) convert  $127\text{ K}$  to  $^\circ\text{C}$ .
2. A gas has an initial volume of  $2.75\text{ L}$  at a temperature of  $285\text{ K}$ . If the temperature changes to  $380\text{ K}$ , what is the new volume of the gas if the pressure is unchanged?
3. Gas can often be cooled by compressing it while keeping the pressure constant. If I have  $45.0\text{ L}$  of gas at room temperature ( $22^\circ\text{C}$ ) and I compress it so that the final volume is  $0.50\text{ L}$ , what is the final temperature of the gas if the pressure is constant?
4. The volume of a gas is  $2.5\text{ L}$  when the pressure is at standard pressure ( $101.325\text{ kPa}$ ). What is the volume of the gas if the pressure decreases to  $85\text{ kPa}$  and the temperature remains unchanged?
5. A  $5.0\text{ L}$  container of gas experiences a temperature change so that the final temperature is 4 times the initial temperature. What is the size of the container after the temperature change? (Assume constant pressure.)
6. At  $45^\circ\text{C}$  the volume of a certain gas is  $27.5\text{ L}$  and the pressure is  $210\text{ kPa}$ . What is the volume of the gas at standard temperature ( $273\text{ K}$ ) and  $310\text{ kPa}$  of pressure?
7. The pressure of a sample of gas was  $97.8\text{ kPa}$  and the volume of the gas was  $3.75\text{ L}$ . If the gas occupied a container with a volume of  $8.00\text{ L}$ , what would the pressure in the container be?
8. Isothermal expansion refers to allowing a gas to expand while keeping the temperature constant. This is one means to simulate a vacuum. If a gas originally at  $97\text{ kPa}$  is allowed to expand from  $0.25\text{ L}$  to  $182\text{ L}$ , what is the pressure of the gas?
9. A gas is initially at a pressure of  $225\text{ kPa}$  and a temperature of  $245\text{ K}$  in a container that is  $4.5\text{ L}$ . If the gas is compressed to a volume of  $2.1\text{ L}$  and the temperature changes to  $275\text{ K}$ , what is the new pressure?

# Gas Laws Practice Skill Practice 35 Answers

**S Ben Porath**



## **Gas Laws Practice Skill Practice 35 Answers:**

## Reviewing **Gas Laws Practice Skill Practice 35 Answers**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Gas Laws Practice Skill Practice 35 Answers**," an enthralling opus penned by a very acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve in to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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