

7. If you have 17 liters of gas at a temperature of 67 °C and a pressure of 88.89 atm, what will be the pressure of the gas if you raise the temperature to 94 °C and decrease the volume to 12 liters?

$$P_2 = \frac{P_1 V_1 T_2}{V_2 T_1} = \frac{88.89 \text{ atm} (17 \text{ L}) 367 \text{ K}}{(12 \text{ L}) 340 \text{ K}}$$

$$P_2 = \quad \text{atm}$$

8. I have an unknown volume of gas at a pressure of 0.5 atm and a temperature of 325 K. If I raise the pressure to 1.2 atm, decrease the temperature to 320 K, and measure the final volume to be 48 liters, what was the initial volume of the gas?

$$V_1 = \frac{P_2 V_2 T_1}{T_2 P_1} = \frac{1.2 \text{ atm} (48 \text{ L}) 325 \text{ K}}{320 \text{ K} (0.5 \text{ atm})}$$

$$V_1 = \quad \text{L}$$

9. If you have 21 liters of gas held at a pressure of 78 atm and a temperature of 900 K, what will be the volume of gas if you decrease the pressure to 45 atm and decrease the temperature to 750 K?

$$V_2 = \frac{78 \text{ atm} (21 \text{ L}) 750 \text{ K}}{900 \text{ K} (45 \text{ atm})}$$

$$V_2 = \quad \text{L}$$

10. If I have 2.9 L of gas at a pressure of 5 atm and a temperature of 50 °C, what will the temperature of the gas be if I decrease the volume of the gas to 2.4 L and decrease the pressure to 3 atm?

$$1/T_2 = \frac{P_1 V_1}{T_1 P_2 V_2} = \frac{(2.9 \text{ L}) 5 \text{ atm}}{323 \text{ K} (3 \text{ atm}) 2.4 \text{ L}}$$

$$T_2 = 160 \text{ K}$$

11. You have an unknown volume of gas held at a temperature of 115 K in a container with a pressure of 60 atm. If by increasing the temperature to 225 K and decreasing the pressure to 30 atm causes the volume of the gas to be 29 L, how many liters of gas did you start with?

$$V_1 = \frac{P_2 V_2 T_1}{T_2 P_1} = \frac{115 \text{ K} (29 \text{ L}) 30 \text{ atm}}{(225 \text{ K}) 60 \text{ atm}}$$

$$V_1 = \quad \text{L}$$

Gas Laws Practice Calculations Key

Ying Liu

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Pp/Chemistry BarCharts, Inc.,2008-06-18

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Dr. Salm's Chemistry Problem Drill Book Robert J. Salm,Robert J. Salm Robert J. Salm Ph.D.,2005 This drill book contains many common problem types that are asked in General Chemistry classes in High School and College This work will give you practice with the major problem types as you prepare for finals and standardized tests

Thermodynamics Problem Solving in Physical Chemistry Kathleen E. Murphy,2020-03-23

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Thermofluids Michael Horsley,Sherwin Keith,1996-12-07

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