

Mixed Gas Law Worksheet

Answers available at end

Name: _____ (show your work) period: _____

Convert:

$$100 \text{ mmHg} = \text{_____ kPa}$$

$$300 \text{ mmHg} = \text{_____ atm}$$

$$120 \text{ atm} = \text{_____ mmHg}$$

$$300 \text{ torr} = \text{_____ kPa}$$

$$25 \text{ kPa} = \text{_____ mmHg}$$

$$25 \text{ mmHg} = \text{_____ atm}$$

$$300^\circ\text{C} = \text{_____ K}$$

$$150\text{K} = \text{_____ }^\circ\text{C}$$

$$5000^\circ\text{C} = \text{_____ K}$$

$$300\text{K} = \text{_____ }^\circ\text{C}$$

$$-255^\circ\text{C} = \text{_____ K}$$

$$3500\text{K} = \text{_____ }^\circ\text{C}$$

$$200^\circ\text{C} = \text{_____ }^\circ\text{F}$$

$$95^\circ\text{F} = \text{_____ }^\circ\text{C} = \text{_____ K}$$

$$45^\circ\text{F} = \text{_____ }^\circ\text{C} = \text{_____ K}$$

- 1) A 70. liter sample of gas initially at 280°C is allowed to cool at constant pressure, what will the new volume be at 0°C ?

$$P_1 = \text{_____} \quad P_2 = \text{_____}$$

$$V_1 = \text{_____} \quad V_2 = \text{_____}$$

$$T_1 = \text{_____} \quad T_2 = \text{_____}$$

- 2) A 4.0 liter sample of gas initially at 25°C is allowed to cool at constant pressure, what will the new volume be at -5°C ?

$$P_1 = \text{_____} \quad P_2 = \text{_____}$$

$$V_1 = \text{_____} \quad V_2 = \text{_____}$$

$$T_1 = \text{_____} \quad T_2 = \text{_____}$$

- 3) A buoyancy vest (used in scuba diving) is filled to a volume of 2.3 liters at a pressure of 1 atmosphere (at the surface where the temperature is a balmy 300K), what will the volume of air be in the vest when the diver dives to a depth of about 90 feet where the temperature is 17°C and the pressure is 4 atmospheres?

$$P_1 = \text{_____} \quad P_2 = \text{_____}$$

Gas Laws And Scuba Diving Worksheet Answers

William Maurice Dehn



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