

1. The theory that explains the behavior of gases at the molecular level is called the _____ theory. _____ reflects its kinetic assumptions about a theoretical gas which are reflected in some _____.
2. Gases behave most free-like gas behavior under conditions of very low _____ and very high _____.
 The molecules of an ideal gas display the _____ as _____ for equilibrium.
 Under ordinary conditions, no ideal gas behaves exactly as _____
 which reflects aspects of the gas and its _____.
 Ideal gas particles move in _____ lines until they collide with each other or with the walls of their container.
 The collisions between the molecules of an ideal gas are completely _____.
 The average kinetic energy of the molecules of an ideal gas is _____ proportional to the _____ temperature of the gas.
3. A gas exerts pressure on the walls of its container because gas molecules _____ with the walls of the container. So, the pressure exerted by a gas depends on how fast the:
 - a) molecules are moving
 - b) _____
4. To measure gas pressure an instrument called a _____ is used.
5. The earth's atmosphere has weight, which creates _____.
6. The instrument used to measure atmospheric pressure is the _____.

Gas Law Note Taking Guide 901 Key

Chao Zhang

A red circular graphic with a gradient, appearing as a partial circle or a stylized 'C' shape, located to the right of the author's name.

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