

Gas Laws HOMEWORK

Lesson 1 Homework

1. Define pressure. *Force that a gas exerts on its container*

2. Convert the following:

a. 750 mmHg to atm

b. 550 mmHg to atm

c. 2.5 atm to torr

d. 2.1×10^3 Pa to atm

e. 300 Pa to atm

$$750 \text{ mmHg} \times (1 \text{ atm} / 760 \text{ mmHg}) = 0.9868 \text{ atm}$$

$$2.5 \text{ atm} \times (760 \text{ mmHg} / 1 \text{ atm}) = 1900 \text{ mmHg}$$

$$2013 \text{ Pa} \times (1 \text{ atm} / 101325 \text{ Pa}) = 0.01986 \text{ atm}$$

$$1700 \text{ mmHg} \times (133 \text{ Pa} / 1 \text{ mmHg}) = 226100 \text{ Pa}$$

$$2.2 \times 10^3 \text{ Pa} \times (1 \text{ atm} / 101325 \text{ Pa}) = 0.0217 \text{ atm}$$

3. Calculate the atmospheric pressure in kPa on a day when the pressure is 750 mm.

$$750 \text{ mm} \times (101325 \text{ Pa} / 760 \text{ mm}) = 100000 \text{ Pa} \rightarrow 100 \text{ kPa}$$

4. Calculate the pressure of a sample of gas (in mmHg) inside an open-end manometer if the mercury column is 35 mm higher on the side of the closed tube than on the atmosphere. The atmospheric pressure is 760 mm. Draw a picture to illustrate your answer.

$$760 \text{ mm} + (1700 \text{ mmHg} / 133 \text{ mm}) = 760 \text{ mmHg}$$

1700 mmHg is 12.8 mm Hg, so the atmospheric pressure, then, is 760 mmHg + 12.8 mmHg.

760 mmHg + 12.8 = 772.8 mmHg
Draw picture showing a manometer with the height of the mercury 35 mm above the atmospheric level of 760 mm.

5. If a gas at a pressure of 150 mm in a 1.0 L container has its pressure increased to 400 mm, what is the new volume of the container? (Assume temperature is constant).

$$P_1 V_1 = P_2 V_2$$

$$(150 \text{ mm} \times 1.0 \text{ L}) = 400 \text{ mm} \times V_2$$

$$V_2 = 0.375 \text{ L} \rightarrow 0.38 \text{ L}$$

6. What is the new pressure if a 5.00 closed container with a pressure of 2.5 atm is increased to 8.0 L?

$$P_1 V_1 = P_2 V_2$$

$$(2.5 \text{ atm} \times 5.00 \text{ L}) = P_2 \times 8.0 \text{ L}$$

$$P_2 = 1.56 \text{ atm}$$

7. When 0.5 L of nitrogen and 0.5 L of sulfur dioxide at 0°C and 1.0 atm, are pumped into an evacuated 1.0 L rigid container, what is the final pressure in the container at 0°C? *Since temperature doesn't change, you can leave it out of the equation.*

So, final pressure of Nitrogen:

$$P_1 V_1 = P_2 V_2$$

$$(1 \text{ atm} \times 0.5 \text{ L}) = P_2 \times 1.0 \text{ L}$$

$$P_2 = 0.5 \text{ atm}$$

So, final pressure of Sulfur dioxide:

$$P_1 V_1 = P_2 V_2$$

$$(1 \text{ atm} \times 0.5 \text{ L}) = P_2 \times 1.0 \text{ L}$$

$$P_2 = 0.5 \text{ atm}$$

Total Pressure = $P_1 + P_2 + P_3$

$$1.0 = 0.5 + 0.5 \rightarrow 1.0 \text{ atm}$$

Assume both containers are 1.0 L each:

$$P_1 V_1 = P_2 V_2$$

$$(1 \text{ atm} \times 1.0 \text{ L}) = P_2 \times 1.0 \text{ L}$$

$$P_2 = 1.0 \text{ atm}$$

Then, total pressure volume and moles are directly proportional!

8. A sealed flask contains a gas with pressure P. If the number of moles of the gas is doubled, the absolute temperature is doubled and the volume of the container is quadrupled, what is the new pressure in terms of P?

Doubling the moles doubles the pressure.

$$P_1 V_1 = P_2 V_2$$

Quadrupling the volume cuts the pressure in 1/4.

$$P_1 V_1 = P_2 V_2$$

moles do not really change in total. P=2

9. If 4 moles of gas are added to a container that already holds 1 mole of gas, how will the pressure change within the container?

$$P_1 V_1 = P_2 V_2$$

Going from 1 mole to 5 moles results in an increase of pressure by 4 times.

$$P_1 V_1 = P_2 V_2$$

10. If a balloon containing 1000 L of gas at 40.°C and 760 mmHg rises to an altitude where the pressure is 500 mmHg and the temperature is 10.0°C, what is the new volume of the balloon (in L)?

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

$$(760 \text{ mmHg} \times 1000 \text{ L}) / 313 \text{ K} = (500 \text{ mmHg} \times V_2) / 273 \text{ K}$$

$$V_2 = 511.76 \text{ L} \rightarrow 512 \text{ L}$$

11. A sample of nitrogen gas has a pressure of 87.5 mmHg in a 500. mL flask. What is the pressure of this gas sample when it is transferred to a 225.0 mL flask at the same temperature?

$$P_1 V_1 = P_2 V_2$$

$$(87.5 \text{ mmHg} \times 500 \text{ mL}) = P_2 \times 225 \text{ mL}$$

$$P_2 = 194.44 \text{ mmHg}$$

12. You have 4.5 L of NO_2 at a temperature of 23.0°C. What volume would the NO_2 occupy at 47°C? (Assume the pressure is constant).

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

$$4.5 \text{ L} / (23.0 + 273 \text{ K}) = V_2 / (47 + 273 \text{ K})$$

$$V_2 = 6.88 \text{ L} \rightarrow 6.9 \text{ L}$$

13. You have a sample of a gas in a flask with a volume of 240 mL. At 35.0°C the pressure of the gas is 580 mmHg. If you decrease the temperature to -5.0°C, what is the gas pressure at the lower temperature? *Since volume doesn't change, you can leave it out of the equation.*

$$P_1 / T_1 = P_2 / T_2$$

$$(580 \text{ mmHg} / 308 \text{ K}) = (P_2 / 268 \text{ K})$$

$$P_2 = 580 \text{ mmHg} \rightarrow 580 \text{ mmHg}$$

Gas Laws Homework Answer Key

Iris Stovall, Donald J. DeCoste



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1. Understanding the eBook Gas Laws Homework Answer Key
 - The Rise of Digital Reading Gas Laws Homework Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Gas Laws Homework Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gas Laws Homework Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Laws Homework Answer Key
 - Personalized Recommendations
 - Gas Laws Homework Answer Key User Reviews and Ratings
 - Gas Laws Homework Answer Key and Bestseller Lists
5. Accessing Gas Laws Homework Answer Key Free and Paid eBooks
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 - Gas Laws Homework Answer Key Compatibility with Devices
 - Gas Laws Homework Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gas Laws Homework Answer Key
 - Highlighting and Note-Taking Gas Laws Homework Answer Key
 - Interactive Elements Gas Laws Homework Answer Key
- 8. Staying Engaged with Gas Laws Homework Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gas Laws Homework Answer Key
- 9. Balancing eBooks and Physical Books Gas Laws Homework Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gas Laws Homework Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Gas Laws Homework Answer Key
 - Setting Reading Goals Gas Laws Homework Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gas Laws Homework Answer Key
 - Fact-Checking eBook Content of Gas Laws Homework Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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