

2

MATTER AND CHANGE

SECTION 2.1 PROPERTIES OF MATTER (pages 39–42)

This section helps you distinguish extensive from intensive properties and identify substances by their properties. It teaches you how to differentiate the three states of matter. It also defines a physical property and lists examples of physical properties and physical changes.

► Describing Matter (page 39)

1. The _____ mass _____ of an object is a measure of the amount of matter the object contains.

2. How does an extensive property differ from an intensive property?

An extensive property depends on the amount of matter, while an intensive property depends on the type of matter.

► Identifying Substances (page 40)

3. Matter that has a uniform and definite composition is called a _____ substance _____.

4. Is the following sentence true or false? All samples of a substance have different physical properties. _____ false _____

5. A physical property is a quality or condition of a substance that can be _____ observed _____ or _____ measured _____ without changing the substance's composition.

6. Circle the letter of the term that is NOT a physical property.

- a. hardness e. boiling point
b. color d. melting

7. Look at Table 2.1 on page 40. What is the melting point of bromine? _____ -7°C _____

8. Look at Table 2.1 on page 40. Circle the letter of the substance that is a yellow solid and melts at 115°C.

- a. sulfur
b. chlorine
c. gold
d. copper

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Table of Contents Matter And Change Guide

1. Understanding the eBook Matter And Change Guide
 - The Rise of Digital Reading Matter And Change Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Matter And Change Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Matter And Change Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Matter And Change Guide
 - Personalized Recommendations
 - Matter And Change Guide User Reviews and Ratings
 - Matter And Change Guide and Bestseller Lists
5. Accessing Matter And Change Guide Free and Paid eBooks

- Matter And Change Guide Public Domain eBooks
- Matter And Change Guide eBook Subscription Services
- Matter And Change Guide Budget-Friendly Options
- 6. Navigating Matter And Change Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Matter And Change Guide Compatibility with Devices
 - Matter And Change Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Matter And Change Guide
 - Highlighting and Note-Taking Matter And Change Guide
 - Interactive Elements Matter And Change Guide
- 8. Staying Engaged with Matter And Change Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Matter And Change Guide
- 9. Balancing eBooks and Physical Books Matter And Change Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Matter And Change Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Matter And Change Guide
 - Setting Reading Goals Matter And Change Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Matter And Change Guide
 - Fact-Checking eBook Content of Matter And Change Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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