

Model 2 – Prefixes and Suffixes

Prefix	Numerical Value
mono-	1
di-	2
tri-	3
tetra-	4
penta-	5
hexa-	6
hepta-	7
octa-	8
nona-	9
deca-	10

Molecular Formula	Name of Compound
BCl_3	Boron trichloride
SF_6	Sulfur hexafluoride
IF_7	Iodine heptafluoride
NI_3	Nitrogen triiodide
N_2O_4	Dinitrogen tetroxide
Cl_2O	Dichlorine monoxide
P_4O_{10}	Tetraphosphorus decoxide
B_5H_9	Pentaboron nonahydride
Br_2O_8	Tribromine octoxide
ClF	Chlorine monofluoride

- Examine the prefixes in Model 2. Fill in the numerical value that corresponds to each prefix.
- What suffix (ending) do all the compound names in Model 2 have in common?

They all end in "ide".

- Carefully examine the names of the compounds in Model 2. When is a prefix NOT used in front of the name of an element?

*When the first element is one. (Examples: IF_7 - iodine heptafluoride)
 NI_3 - nitrogen triiodide)*

- Consider the compound NO.

- Which element, nitrogen or oxygen, would require a prefix in the molecule name? Explain your answer.

Oxygen would require a prefix of "mono" because it is the second element and there is one oxygen. Nitrogen does not require a prefix because it is the first element and there is one nitrogen.

- Name the molecule NO.

*Nitrogen monoxide (Notice, the "o" of mono was dropped)
 ↳ not monoxide*

- Find two compounds in Model 2 that contain a subscript of "4" in their molecular formula.

- List the formulas and names for the two compounds.

N_2O_4 - dinitrogen tetroxide || P_4O_{10} - Tetraphosphorus decoxide

- What is different about the spelling of the prefix meaning "four" in these two names?

The "a" has been dropped from the end of the prefix tetra.

If8766 Answer Key Naming Molecular Compounds

R Barnett



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