

Name _____ date _____

Ideal Gas Law Worksheet Review

$$PV=nRT$$

mass \longleftrightarrow mole

$$R = \begin{array}{l} 0.0821 \text{ L}\cdot\text{atm/mol}\cdot\text{K} \\ 62.4 \text{ L}\cdot\text{mm Hg/mol}\cdot\text{K} \\ 8.314 \text{ L}\cdot\text{kPa/mol}\cdot\text{K} \end{array}$$

Standard pressure and temperature (STP): 1 atm and 273 K

1. How many moles of a gas at 100°C does it take to fill a 1.00 L flask to a pressure of 1.50 atm?
2. A camping stove holds 3000 g of liquid C_3H_8 . How large of a container would be needed to hold the gas at a temperature of at 25°C and a pressure of 3.0 atm? (hint: need to convert to moles first)
3. What volume would be occupied by 3.5 mol of oxygen gas at a pressure of 768 mm Hg and a temperature of 25°C?
4. A ten-liter gas cylinder contains 55 moles nitrogen. What pressure, in kPa, is exerted by the nitrogen at 25°C?
5. A drum used to transport crude oil has a volume of 162 L. How many grams of water, as steam, are required to fill the drum at 1.20 atm and 100°C?
6. How many moles of air are there in a 125 mL flask if the pressure is 755 mm Hg and the temperature is 20°C?

Ideal Gas Law Worksheet With Answers

Christian G. Meyer

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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