#### Ionic Compounds & Properties

Ionic compounds typically exist as non-volatile solids due to the strong ionic bonds holding the lattice together. This means that the ions are stuck in their positions in the lattice and are not free to move. However, when dissolved in water, the water molecules surround the ions and separate them from each other; this means that the ions are now free to move. An electrical current is simply the flow of charge, so ionic compounds can conduct electricity when dissolved or molten.

# Metallic Bonding & Alloys

If we consider elemental metals or alloys, there is only one type of atom - metals. Metals have low electronegativity, but also low differences in electronegativity - This is how a metallic bond is defined. The bonding in metals is between the positive nuclei and the sea of delocalized electrons. This means that the strength of the bond may be determined by the charge density of the cation and the number of electrons that are delocalized over the lattice. Alloys are typically mixtures of two or more metals (and possibly carbon) and have properties proportional to the ratio of the metals. The properties depend on the properties and lattice structure of the constituent atoms.

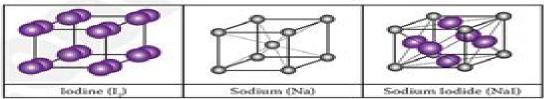
# Melting Point

Melting points are a simple way of estimating the relative strength of bonds between whatever is making up the lattice. Typically giant structures have high melting points because of the strong forces holding the lattice together. Even so, we can use melting points to describe the trends in the strength of the various types of bonds.

The melting point is the temperature that gives the particles enough kinetic energy so that they can overcome the attractive forces acting in the lattice of the solid crystal.

Melting points are a measure of the attractive forces that hold the lattice together.

Melting does not decompose a compound into its elements - it is only a change of state! Many students confuse the overcoming of intermolecular forces with decomposition. In most cases, a covalent molecule will change state by overcoming intermolecular forces, well before it decompose by breaking the stronger covalent bonds that hold the molecule together. Consider Iodine, I,, the energy of vaporization is 41 kJ mol\* vs. the bond strength of 151 kJ mol\*.



## Melting points of Group I - The alkali metals

Before we discuss the melting point, we must recall the type of bonding - in this case metallic.

Metallic bonding is the attractive force between the lattice of positive ions, and the "sea" of delocalized electrons. There are three factors affecting the strength of the attraction:

The charge on the ion (eg. +1, +2, +3 for Na, Mg, and Al respectively)

The size of the ion. – as the ions increases in size, the charge density decreases.

3) - The number of delocalized electrons (the opposite charge as the ion)

Know it!

# **Ib Chemistry Hl 2013 May**

**P N Ravindran** 

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knowledge We tried to cover as wide a spectrum as possible and succeeded in gathering an expert and enthusiastic team of speakers but there are some inevitable omissions In particular and with hindsight our discussions might have been enriched by more detailed coverage of general aspects of chemical catalysis but readers requiring this background should find adequate references herein **Cumulated Index Medicus** ,1996 Comprehensive Toxicology, 2017-12-01 Comprehensive Toxicology Third Edition Fifteen Volume Set discusses chemical effects on biological systems with a focus on understanding the mechanisms by which chemicals induce adverse health effects Organized by organ system this comprehensive reference work addresses the toxicological effects of chemicals on the immune system the hematopoietic system cardiovascular system respiratory system hepatic toxicology renal toxicology gastrointestinal toxicology reproductive and endocrine toxicology neuro and behavioral toxicology developmental toxicology and carcinogenesis also including critical sections that cover the general principles of toxicology cellular and molecular toxicology biotransformation and toxicology testing and evaluation Each section is examined in state of the art chapters written by domain experts providing key information to support the investigations of researchers across the medical veterinary food environment and chemical research industries and national and international regulatory agencies Thoroughly revised and expanded to 15 volumes that include the latest advances in research and uniquely organized by organ system for ease of reference and diagnosis this new edition is an essential reference for researchers of toxicology Organized to cover both the fundamental principles of toxicology and unique aspects of major organ systems Thoroughly revised to include the latest advances in the toxicological effects of chemicals on the immune system Features additional coverage throughout and a new volume on toxicology of the hematopoietic system Presents in depth comprehensive coverage from an international author base of domain experts

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