

1. How do the bond angles in  $\text{CH}_4$ ,  $\text{NH}_3$  and  $\text{H}_2\text{O}$  compare?

- A.  $\text{CH}_4 = \text{NH}_3 = \text{H}_2\text{O}$
- B.  $\text{CH}_4 < \text{NH}_3 < \text{H}_2\text{O}$
- C.  $\text{NH}_3 < \text{CH}_4 < \text{H}_2\text{O}$
- D.  $\text{H}_2\text{O} < \text{NH}_3 < \text{CH}_4$

2.

The electronegativities of four different elements are given below (the letters are not their chemical symbols).

Element	W	X	Y	Z
Electronegativity	0.9	1.2	3.4	4.0

Based on this information which statement is correct?

- A. W is a non-metal.
- B. W and X form an ionic compound.
- C. Y is a metal.
- D. Y and Z form a covalent compound.

3.

Which properties of the alkali metals decrease going down group 1?

- A. First ionization energy and reactivity
- B. Melting point and atomic radius
- C. Reactivity and electronegativity
- D. First ionization energy and melting point

# Ib Chemistry Exam 2013 Answers

**Wendy Heydorn, Susan Jesudason**



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