Name of the student

Date – 14.09.2022 (Wednesday) Section

#### GENERAL INSTRUCTIONS:

- All questions are compulsory.
- Question nos. 1 to 21 are MCQs with one correct option and carrying 1 mark each.
- Question nos. 22 to 24 are objective/short answer type questions.
- Question nos. 25 to 28 are of 2 marks each.
- Question nos. 29 to 32 is of 3 marks each.
- Question nos. 33 to 35 is of 5 marks each.

#### SECTION - A

#### Q. 1 to 21 is of 1 mark each.

- Q.1 Which of the following aqueous solution should have highest boiling point?
  - a) 1 M NaOH
- b) 1 M Na<sub>2</sub>SO<sub>4</sub>
- c) 1M NH<sub>4</sub>NO<sub>3</sub>
- d) 1 M KNO<sub>3</sub>

- Q.2 Which of the following statement is false?
  - a) Two different solution of sucrose of same molality prepare in different solvent will have the same depression in freezing point.
  - b) The osmotic pressure of a solution is given by the equation  $\pi$ =CRT. (where C is the molarity of the solution)
  - c) Decreasing order of osmotic pressure for 0.01M aqueous solution of Barium Chloride, Potassium chloride, acetic acid and sucrose is BaCl<sub>2</sub> > KCl > CH<sub>3</sub>COOH > Sucrose
  - d) According to Raoult's Law, the vapour pressure exeterd by a volatile component of solution is directly proportional to its molar fraction in the solution.
- Q.3 We have three aqueous solution of NaCl labelled as 'A', 'B' and 'C' with concentration 0.1 M, 0.01 M and 0.001 M respectively. The value of Van't Hoff factor for these solution will be in the order
  - a) ia < in < ic
- a)  $i_A > i_B > i_C$
- a)  $i_A = i_B = i_C$

- a) iA < iB > ic
- Q.4 On the basis of the information given below mark the correct options -
  - In bromoethane and chloroethane mixture intermolecular interaction of A-A and B-B type are nearly same as A-B type interaction.
  - ii) In the ethanol and acetone mixture A-A or B-B type intermolecular interactions are stronger than A-B type interaction.

# **Hsc Half Yearly Past Papers And Chemistry**

**JS Bruner** 

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#### **Table of Contents Hsc Half Yearly Past Papers And Chemistry**

- 1. Understanding the eBook Hsc Half Yearly Past Papers And Chemistry
  - The Rise of Digital Reading Hsc Half Yearly Past Papers And Chemistry
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Hsc Half Yearly Past Papers And Chemistry
  - Exploring Different Genres
  - o Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Hsc Half Yearly Past Papers And Chemistry
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Hsc Half Yearly Past Papers And Chemistry
  - Personalized Recommendations
  - $\circ\,$  Hsc Half Yearly Past Papers And Chemistry User Reviews and Ratings
  - Hsc Half Yearly Past Papers And Chemistry and Bestseller Lists
- 5. Accessing Hsc Half Yearly Past Papers And Chemistry Free and Paid eBooks
  - Hsc Half Yearly Past Papers And Chemistry Public Domain eBooks
  - Hsc Half Yearly Past Papers And Chemistry eBook Subscription Services
  - Hsc Half Yearly Past Papers And Chemistry Budget-Friendly Options

- 6. Navigating Hsc Half Yearly Past Papers And Chemistry eBook Formats
  - o ePub, PDF, MOBI, and More
  - Hsc Half Yearly Past Papers And Chemistry Compatibility with Devices
  - Hsc Half Yearly Past Papers And Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Hsc Half Yearly Past Papers And Chemistry
  - Highlighting and Note-Taking Hsc Half Yearly Past Papers And Chemistry
  - Interactive Elements Hsc Half Yearly Past Papers And Chemistry
- 8. Staying Engaged with Hsc Half Yearly Past Papers And Chemistry
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Hsc Half Yearly Past Papers And Chemistry
- 9. Balancing eBooks and Physical Books Hsc Half Yearly Past Papers And Chemistry
  - $\circ$  Benefits of a Digital Library
  - Creating a Diverse Reading Collection Hsc Half Yearly Past Papers And Chemistry
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Hsc Half Yearly Past Papers And Chemistry
  - Setting Reading Goals Hsc Half Yearly Past Papers And Chemistry
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Hsc Half Yearly Past Papers And Chemistry
  - Fact-Checking eBook Content of Hsc Half Yearly Past Papers And Chemistry
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements

• Interactive and Gamified eBooks

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