

Ideal Gas Law Practice Worksheet

Solve the following problems using the ideal gas law:

- 1) How many moles of gas does it take to occupy 120 liters at a pressure of 2.3 atmospheres (1.00 atm = 101.3 kPa) and a temperature of 340 K?
- 2) If I have a 50.0 liter container that holds 45.0 moles of gas at a temperature of 200° C, what is the pressure inside the container?
- 3) It is not safe to put aerosol canisters in a campfire, because the pressure inside the canisters gets very high and they can explode. If I have a 1.0 liter canister that holds 2.0 moles of gas, and the campfire temperature is 1400° C, what is the pressure inside the canister?
- 4) How many moles of gas are in a 30 liter scuba canister if the temperature of the canister is 300 K and the pressure is 20.0 atmospheres?
- 5) I have a balloon that can hold 10.0 liters of air. If I blow up this balloon with 0.470 moles of helium gas at a pressure of 120 kPa, at what temperature (°C) is the balloon?
- 6) What is the molar mass of a gas if 10.0 grams of it occupy 2.87 dm³ at a temperature of 25.0°C and 196 kPa?
- 7) What is the density (in g/L) of Argon gas when the temperature is 35.°C and the pressure is 1.50 atm?

Gas Laws Practice Problems With Solutions

Victor M. Corman



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