

Mark Scheme

Q1.

Question Number	Acceptable Answer	Additional Guidance	Mark
(a) (i)	$\text{CaCO}_3(\text{s}) + 2\text{HCl}(\text{aq}) \rightarrow \text{CaCl}_2(\text{aq}) + \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$ <p>Balanced equation (1)</p> <p>State symbols (1)</p>	<p>Accept $\text{CaCO}_3(\text{s}) + 2\text{H}^+(\text{aq}) \rightarrow \text{Ca}^{2+}(\text{aq}) + \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$</p> <p>2nd mark dependent on first or near miss.</p> <p>Reject $\text{H}_2\text{CO}_3(\text{aq})$ in equation, but allow state symbol mark if otherwise correct.</p>	(2)
Question Number	Acceptable Answer	Additional Guidance	Mark
(a) (ii)	Finds molar mass of calcium carbonate	<p>Example of calculation $\text{Mr of calcium carbonate} = 40.1 + 12 + (16 \times 3) = 100.1 \text{ (g mol}^{-1}\text{)}$ Allow $= 40 + 12 + (16 \times 3) = 100 \text{ (g mol}^{-1}\text{)}$ Accept answer with no working</p>	(1)

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Table of Contents May 22012 Chemistry Mark Scheme Edexcel

1. Understanding the eBook May 22012 Chemistry Mark Scheme Edexcel
 - The Rise of Digital Reading May 22012 Chemistry Mark Scheme Edexcel
 - Advantages of eBooks Over Traditional Books
2. Identifying May 22012 Chemistry Mark Scheme Edexcel
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an May 22012 Chemistry Mark Scheme Edexcel
 - User-Friendly Interface
4. Exploring eBook Recommendations from May 22012 Chemistry Mark Scheme Edexcel
 - Personalized Recommendations
 - May 22012 Chemistry Mark Scheme Edexcel User Reviews and Ratings
 - May 22012 Chemistry Mark Scheme Edexcel and Bestseller Lists
5. Accessing May 22012 Chemistry Mark Scheme Edexcel Free and Paid eBooks
 - May 22012 Chemistry Mark Scheme Edexcel Public Domain eBooks
 - May 22012 Chemistry Mark Scheme Edexcel eBook Subscription Services
 - May 22012 Chemistry Mark Scheme Edexcel Budget-Friendly Options

6. Navigating May 22012 Chemistry Mark Scheme Edexcel eBook Formats
 - ePub, PDF, MOBI, and More
 - May 22012 Chemistry Mark Scheme Edexcel Compatibility with Devices
 - May 22012 Chemistry Mark Scheme Edexcel Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of May 22012 Chemistry Mark Scheme Edexcel
 - Highlighting and Note-Taking May 22012 Chemistry Mark Scheme Edexcel
 - Interactive Elements May 22012 Chemistry Mark Scheme Edexcel
8. Staying Engaged with May 22012 Chemistry Mark Scheme Edexcel
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers May 22012 Chemistry Mark Scheme Edexcel
9. Balancing eBooks and Physical Books May 22012 Chemistry Mark Scheme Edexcel
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection May 22012 Chemistry Mark Scheme Edexcel
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine May 22012 Chemistry Mark Scheme Edexcel
 - Setting Reading Goals May 22012 Chemistry Mark Scheme Edexcel
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of May 22012 Chemistry Mark Scheme Edexcel
 - Fact-Checking eBook Content of May 22012 Chemistry Mark Scheme Edexcel
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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