

Limiting Reactant Worksheet #2

1. Given the following reaction: (Balance the equation first!)



- If you start with 14.8 g of C_3H_8 and 3.44 g of O_2 , determine the limiting reactant.
- Determine the number of moles of carbon dioxide produced.
- Determine the number of grams of H_2O produced.
- Determine the number of grams of excess reactant left.

2. Given the following equation:



- If 10.0 g of $\text{Al}_2(\text{SO}_4)_3$ is reacted with 10.0 g of NaOH , determine the limiting reactant.
- Determine the number of moles of $\text{Al}(\text{OH})_3$ produced.
- Determine the number of grams of Na_2SO_4 produced.
- Determine the number of grams of excess reactant left over in the reaction.

3. Given the following equation:



- If 25.4 g of Al_2O_3 is reacted with 10.2 g of Fe , determine the limiting reactant.
- Determine the number of moles of Al produced.
- Determine the number of grams of Fe_3O_4 produced.
- Determine the number of grams of excess reactant left over in the reaction.

4. Consider the reaction



- 80.0 grams of I_2O_5 , reacts with 28.0 grams of carbon monoxide, CO . Determine the mass of iodine I_2 , which could be produced?
- If, in the above situation, only 0.160 moles, of iodine, I_2 was produced.
 - What mass of iodine was produced?
 - What percentage yield of iodine was produced?

5. Zinc and sulfur react to form zinc sulfide according to the equation



If 25.0 g of zinc and 30.0 g of sulfur are mixed,

- Which chemical is the limiting reactant?
- How many grams of ZnS will be formed?
- How many grams of the excess reactant will remain after the reaction is over?

6. Which element is in excess when 3.00 grams of Mg is ignited in 2.20 grams of pure oxygen? What mass is in excess? What mass of MgO is formed?

7. How many grams of Al_2S_3 are formed when 5.00 grams of Al is heated with 10.0 grams S ?

Limiting Reagent Answers And Work

Richard Bailey



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Table of Contents Limiting Reagent Answers And Work

1. Understanding the eBook Limiting Reagent Answers And Work
 - The Rise of Digital Reading Limiting Reagent Answers And Work
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reagent Answers And Work
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reagent Answers And Work
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reagent Answers And Work
 - Personalized Recommendations
 - Limiting Reagent Answers And Work User Reviews and Ratings
 - Limiting Reagent Answers And Work and Bestseller Lists
5. Accessing Limiting Reagent Answers And Work Free and Paid eBooks

- Limiting Reagent Answers And Work Public Domain eBooks
- Limiting Reagent Answers And Work eBook Subscription Services
- Limiting Reagent Answers And Work Budget-Friendly Options
- 6. Navigating Limiting Reagent Answers And Work eBook Formats
 - ePub, PDF, MOBI, and More
 - Limiting Reagent Answers And Work Compatibility with Devices
 - Limiting Reagent Answers And Work Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limiting Reagent Answers And Work
 - Highlighting and Note-Taking Limiting Reagent Answers And Work
 - Interactive Elements Limiting Reagent Answers And Work
- 8. Staying Engaged with Limiting Reagent Answers And Work
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reagent Answers And Work
- 9. Balancing eBooks and Physical Books Limiting Reagent Answers And Work
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reagent Answers And Work
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Limiting Reagent Answers And Work
 - Setting Reading Goals Limiting Reagent Answers And Work
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Limiting Reagent Answers And Work
 - Fact-Checking eBook Content of Limiting Reagent Answers And Work
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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