

Limiting Reagents and Percent Yield Worksheet

Problem 1: Limiting Reagent

1. Consider the following reaction: $2\text{Al} + 3\text{H}_2\text{SO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + 3\text{H}_2$
 When 10.0 grams of Aluminum reacts with excess of H_2SO_4 , what are the limiting and excess reagents? How many grams of $\text{Al}_2(\text{SO}_4)_3$ can be produced?



2. Consider the following reaction: $2\text{Fe} + 3\text{Cl}_2 \rightarrow 2\text{FeCl}_3$
 When 20.0 grams of Fe reacts with 20.0 grams of Cl_2 , what are the limiting and excess reagents? How many grams of FeCl_3 can be produced?



3. Consider the following reaction: $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 When 10.0 grams of H_2 reacts with 10.0 grams of O_2 , what are the limiting and excess reagents? How many grams of H_2O can be produced?



4. Consider the following reaction: $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 If 10.0 grams of H_2 reacts with 10.0 grams of O_2 , determine the limiting reagent, excess reagent, and how many grams of H_2O can be produced?



Limiting Reagent And Percent Yield Study Guide

Martha Mackin



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