

Limiting Reactant and Percent Yield Worksheet

Key

1. Calcium hydroxide, used to neutralize acid spills, reacts with hydrochloric acid according to the following unbalanced equation:



- a. If you have spilled 6.3 mol of HCl and put 2.8 mol of Ca(OH)_2 on it, which substance is the limiting reactant?

$$\frac{6.3 \text{ mol HCl}}{2 \text{ mol HCl}} \bigg| \frac{1 \text{ mol CaCl}_2}{1 \text{ mol Ca(OH)}_2} = 3.2 \text{ mol CaCl}_2$$

$$\frac{2.8 \text{ mol Ca(OH)}_2}{1 \text{ mol Ca(OH)}_2} \bigg| \frac{1 \text{ mol CaCl}_2}{1 \text{ mol Ca(OH)}_2} = 2.8 \text{ mol CaCl}_2$$

* Ca(OH)_2 is limiting

- b. How many moles of the excess reactant remain?

$$\frac{2.8 \text{ mol Ca(OH)}_2}{1 \text{ mol Ca(OH)}_2} \bigg| \frac{2 \text{ mol HCl}}{1 \text{ mol Ca(OH)}_2} = 5.6 \text{ mol HCl used}$$

$$6.3 \text{ mol HCl} - 5.6 \text{ mol HCl} = \boxed{0.7 \text{ mol HCl remain}}$$

2. Aluminum oxidizes according to the following unbalanced equation: $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$

- a. Powdered Al (0.048 mol) is placed into a container containing 0.030 mol O_2 . What is limiting reactant?

$$\frac{0.048 \text{ mol Al}}{4 \text{ mol Al}} \bigg| \frac{2 \text{ mol Al}_2\text{O}_3}{3 \text{ mol O}_2} = 0.024 \text{ mol Al}_2\text{O}_3$$

$$\frac{0.030 \text{ mol O}_2}{3 \text{ mol O}_2} \bigg| \frac{2 \text{ mol Al}_2\text{O}_3}{3 \text{ mol O}_2} = 0.020 \text{ mol Al}_2\text{O}_3$$

* O_2 is limiting

- b. How many moles of the excess reactant remain?

$$\frac{0.030 \text{ mol O}_2}{3 \text{ mol O}_2} \bigg| \frac{4 \text{ mol Al}}{3 \text{ mol O}_2} = 0.040 \text{ mol Al used}$$

$$0.048 \text{ mol Al} - 0.040 \text{ mol Al} = \boxed{0.008 \text{ mol Al remain}}$$

3. Heating zinc sulfide in the presence of oxygen yields the following: $2\text{ZnS} + 3\text{O}_2 \rightarrow 2\text{ZnO} + 2\text{SO}_2$

- a. If 1.72 mol of ZnS is heated in the presence of ^{45.6}~~32.0~~ g of O_2 , which is the limiting reactant? (Balance the equation first.)

$$\frac{1.72 \text{ mol ZnS}}{1 \text{ mol ZnS}} \bigg| \frac{2 \text{ mol SO}_2}{1 \text{ mol ZnS}} = 3.44 \text{ mol SO}_2$$

$$\frac{45.6 \text{ g O}_2}{32.0 \text{ g O}_2} \bigg| \frac{1 \text{ mol O}_2}{32.0 \text{ g O}_2} \bigg| \frac{2 \text{ mol SO}_2}{3 \text{ mol O}_2} = 0.950 \text{ mol SO}_2$$

* O_2 is limiting

- b. How many grams of the excess reactant remain?

$$\frac{45.6 \text{ g O}_2}{32.0 \text{ g O}_2} \bigg| \frac{1 \text{ mol O}_2}{32.0 \text{ g O}_2} \bigg| \frac{2 \text{ mol ZnS}}{3 \text{ mol O}_2} \bigg| \frac{97.48 \text{ g ZnS}}{1 \text{ mol ZnS}} = 92.6 \text{ g used}$$

$$\frac{1.72 \text{ mol ZnS}}{1 \text{ mol ZnS}} \bigg| \frac{97.48 \text{ g ZnS}}{1 \text{ mol ZnS}} = 168 \text{ g ZnS (started with)}$$

$$\frac{168 \text{ g ZnS}}{92.6 \text{ g ZnS}} = \boxed{75 \text{ g ZnS remain}}$$

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Limiting Reagent And Percent Yield Answer Key 12 3:

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Table of Contents Limiting Reagent And Percent Yield Answer Key 12 3

1. Understanding the eBook Limiting Reagent And Percent Yield Answer Key 12 3
 - The Rise of Digital Reading Limiting Reagent And Percent Yield Answer Key 12 3
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reagent And Percent Yield Answer Key 12 3
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reagent And Percent Yield Answer Key 12 3
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reagent And Percent Yield Answer Key 12 3
 - Personalized Recommendations
 - Limiting Reagent And Percent Yield Answer Key 12 3 User Reviews and Ratings
 - Limiting Reagent And Percent Yield Answer Key 12 3 and Bestseller Lists

5. Accessing Limiting Reagent And Percent Yield Answer Key 12 3 Free and Paid eBooks
 - Limiting Reagent And Percent Yield Answer Key 12 3 Public Domain eBooks
 - Limiting Reagent And Percent Yield Answer Key 12 3 eBook Subscription Services
 - Limiting Reagent And Percent Yield Answer Key 12 3 Budget-Friendly Options
6. Navigating Limiting Reagent And Percent Yield Answer Key 12 3 eBook Formats
 - ePub, PDF, MOBI, and More
 - Limiting Reagent And Percent Yield Answer Key 12 3 Compatibility with Devices
 - Limiting Reagent And Percent Yield Answer Key 12 3 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limiting Reagent And Percent Yield Answer Key 12 3
 - Highlighting and Note-Taking Limiting Reagent And Percent Yield Answer Key 12 3
 - Interactive Elements Limiting Reagent And Percent Yield Answer Key 12 3
8. Staying Engaged with Limiting Reagent And Percent Yield Answer Key 12 3
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reagent And Percent Yield Answer Key 12 3
9. Balancing eBooks and Physical Books Limiting Reagent And Percent Yield Answer Key 12 3
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reagent And Percent Yield Answer Key 12 3
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Limiting Reagent And Percent Yield Answer Key 12 3
 - Setting Reading Goals Limiting Reagent And Percent Yield Answer Key 12 3
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Limiting Reagent And Percent Yield Answer Key 12 3
 - Fact-Checking eBook Content of Limiting Reagent And Percent Yield Answer Key 12 3
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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