#### Limiting Reactants Worksheet KEY

Answer the following questions, using stoichiometry to convert between moles of one substance and moles of another substance within a reaction and using molar masses to convert between grams of a substance and moles of that substance.

#### Given the following reaction: $C_3H_{12} + 8 O_2 \rightarrow 5 CO_2 + 6 H_2O$

 If 5.25 grams of pentane (C<sub>1</sub>H<sub>12</sub>) are completely reacted with oxygen (O<sub>2</sub>), how many grams of carbon dioxide (CO<sub>2</sub>) will be produced?

5.25 g C<sub>3</sub>H<sub>12</sub> x 
$$\frac{1 \, mol \, C_5 \, H_{12}}{72.151 \, g \, C_5 \, H_{12}} = 0.07276 \, mol \, C_5 H_{12}$$
  
0.07276 mol C<sub>3</sub>H<sub>12</sub> x  $\frac{5 \, mol \, C \, O_2}{1 \, mol \, C_5 \, H_{12}} = 0.3638 \, mol \, CO_2$   
0.3638 mol CO<sub>2</sub> x  $\frac{44.009 \, g \, CO_2}{1 \, mol \, CO_2} = 16.0 \, g \, CO_2$ 

 If 9.75 grams of oxygen (O<sub>2</sub>) are completely reacted with pentane (C<sub>3</sub>H<sub>12</sub>), how many grams of carbon dioxide (CO<sub>2</sub>) will be produced?

9.75 g O<sub>2</sub> x 
$$\frac{1 mol O_2}{31.998 g O_2}$$
 = 0.3047 mol O<sub>2</sub>  
0.3047 mol O<sub>2</sub> x  $\frac{5 mol C O_2}{8 mol O_2}$  = 0.1904 mol CO<sub>2</sub>  
0.1904 mol CO<sub>2</sub> x  $\frac{44.009 g CO_2}{1 mol CO_2}$  = 8.38 g CO<sub>2</sub>

3. Based on your answers to question #1 and 2, if 5.25 grams of pentane are reacted with 9.75 grams of oxygen, what is the limiting reactant and how many grams of CO<sub>2</sub> actually form?

Limiting reactant: O2

8.38 g CO2 actually form

# **Limiting Reactants Key**

Morris Hein, Scott Pattison, Susan Arena, Leo R. Best

## **Limiting Reactants Key:**

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