Limiting Reactants Worksheet KEY

Answer the following questions, using stoichiometry to convert between moles of one substance and moles of another substance within a reaction and using molar masses to convert between grams of a substance and moles of that substance.

Given the following reaction: $C_3H_{12} + 8 O_2 \rightarrow 5 CO_2 + 6 H_2O$

 If 5.25 grams of pentane (C₁H₁₂) are completely reacted with oxygen (O₂), how many grams of carbon dioxide (CO₂) will be produced?

5.25 g C₃H₁₂ x
$$\frac{1 \, mol \, C_5 \, H_{12}}{72.151 \, g \, C_5 \, H_{12}} = 0.07276 \, mol \, C_5 H_{12}$$

0.07276 mol C₃H₁₂ x $\frac{5 \, mol \, C \, O_2}{1 \, mol \, C_5 \, H_{12}} = 0.3638 \, mol \, CO_2$
0.3638 mol CO₂ x $\frac{44.009 \, g \, CO_2}{1 \, mol \, CO_2} = 16.0 \, g \, CO_2$

 If 9.75 grams of oxygen (O₂) are completely reacted with pentane (C₃H₁₂), how many grams of carbon dioxide (CO₂) will be produced?

9.75 g O₂ x
$$\frac{1 mol O_2}{31.998 g O_2}$$
 = 0.3047 mol O₂
0.3047 mol O₂ x $\frac{5 mol C O_2}{8 mol O_2}$ = 0.1904 mol CO₂
0.1904 mol CO₂ x $\frac{44.009 g CO_2}{1 mol CO_2}$ = 8.38 g CO₂

3. Based on your answers to question #1 and 2, if 5.25 grams of pentane are reacted with 9.75 grams of oxygen, what is the limiting reactant and how many grams of CO₂ actually form?

Limiting reactant: O2

8.38 g CO2 actually form

Limiting Reactants C1y Vm Answers

Annelies Wilder-Smith

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