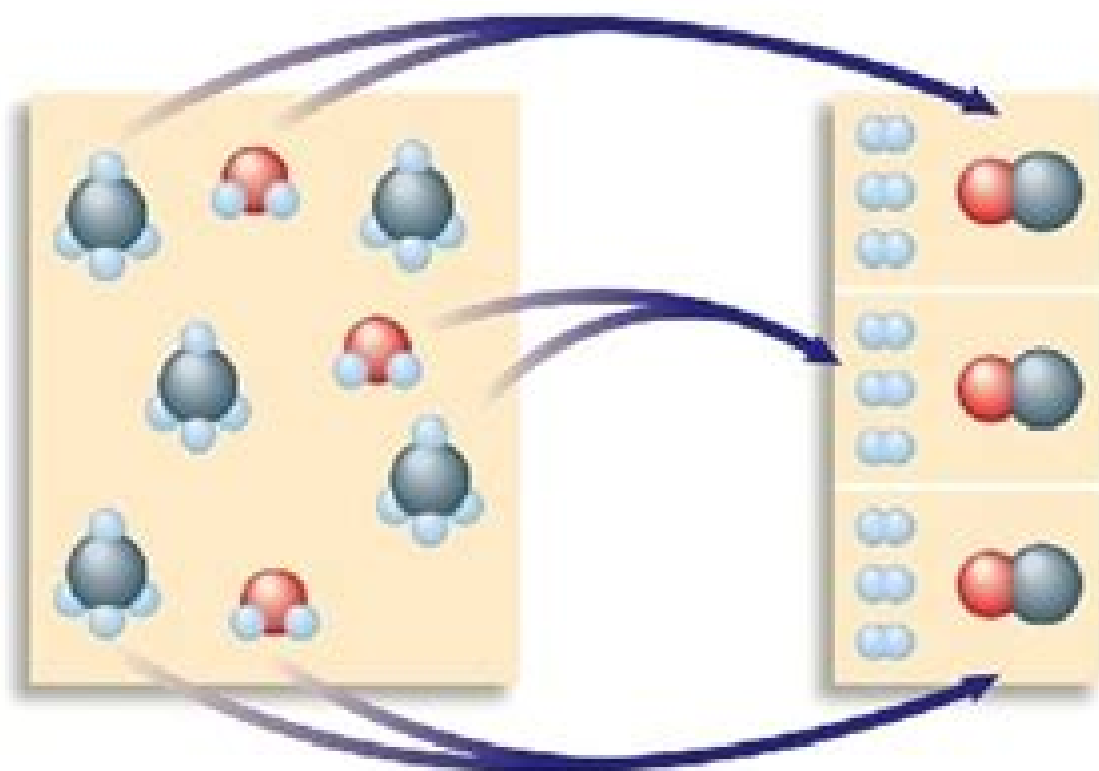


## Limiting Reactants and Percent Yield

## B. Calculations Involving a Limiting Reactant

A mixture of 5  $\text{CH}_4$  and 3  $\text{H}_2\text{O}$  molecules undergoes the reaction  
 $\text{CH}_4(g) + \text{H}_2\text{O}(g) \rightarrow 3\text{H}_2(g) + \text{CO}(g)$ .  
Note that the  $\text{H}_2\text{O}$  molecules are used up first, leaving two  $\text{CH}_4$  molecules unreacted.



# Limiting Reactants And Percent Yield Section 9 3 Notes

**Ibrahim Sikder**



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