

1. What volume of hydrogen gas at STP is produced from the reaction of 50.0g of Mg and 75.0 grams of HCl? How much of the excess reagent is left over (in grams)?



Do a standard mass to mass problem starting with each reactant

50.0 g Mg	$\frac{1 \text{ mol Mg}}{24.31 \text{ g Mg}}$	$\frac{1 \text{ mol H}_2}{1 \text{ mol Mg}}$	$\frac{22.4 \text{ L}}{1 \text{ mol H}_2}$	= 46.1 L H ₂
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75.0 g HCl	$\frac{1 \text{ mol HCl}}{36.46 \text{ g HCl}}$	$\frac{1 \text{ mole H}_2}{2 \text{ mol HCl}}$	$\frac{22.4 \text{ L}}{1 \text{ mol H}_2}$	= 23.0 L H ₂
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HCl is the limiting reactant!!

Limiting Reactant Pogil Key

Lifelique



Limiting Reactant Pogil Key:

Limiting Reactant and Percent Lifelique, 2019 This lesson plan covers analyzing chemical reactions in order to determine limiting reactants and excess reactants including calculating the amount of excess reactant calculate the theoretical yield of a reaction and calculating the percent yield of a reaction

Decoding **Limiting Reactant Pogil Key**: Revealing the Captivating Potential of Verbal Expression

In a time characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its capability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Limiting Reactant Pogil Key**," a mesmerizing literary creation penned by a celebrated wordsmith, readers embark on an enlightening odyssey, unraveling the intricate significance of language and its enduring affect our lives. In this appraisal, we shall explore the book is central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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