

## Calculations need to be based on the limiting reactant.

- Example 2: What is the mass of sulfur dioxide that can be produced by the reaction of 95.6 g carbon disulfide with 100. g oxygen?



$$\frac{95.6\text{g CS}_2}{76.2\text{g CS}_2} \times \frac{1\text{mol CS}_2}{1\text{mol CS}_2} \times \frac{2\text{mol SO}_2}{1\text{mol CS}_2} \times \frac{64.1\text{g SO}_2}{1\text{mol SO}_2} = 161\text{g SO}_2$$

$$\frac{100.\text{g O}_2}{32.0\text{g O}_2} \times \frac{1\text{mol O}_2}{3\text{mol O}_2} \times \frac{2\text{mol SO}_2}{1\text{mol O}_2} \times \frac{64.1\text{g SO}_2}{1\text{mol SO}_2} = 134\text{g SO}_2$$

134 g of SO<sub>2</sub> can be produced in this reaction.

# Limiting Reactant Answers

**David Baud**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located on the right side of the slide, partially overlapping the light blue bar.

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