

Limiting Reactants Practice KEY

1. How many grams of aluminum chloride could be produced from 35.5 g of aluminum and 39.0 g of chlorine gas?



$$35.5 \text{ g Al} \times \frac{1 \text{ mol Al}}{26.982 \text{ g Al}} \times \frac{3 \text{ mol Cl}_2}{2 \text{ mol Al}} \times \frac{70.906 \text{ g Cl}_2}{1 \text{ mol Cl}_2} = 140. \text{ g Cl}_2$$

- ∞ This means 35.5 g Al would react with 140. g Cl₂
- ∞ Only 39.0 g Cl₂ were available, so it is the limiting reactant.
- ∞ Use the mass of Cl₂ in the reaction to determine the mass of the AlCl₃ formed.

$$39.0 \text{ g Cl}_2 = \text{ ? } \text{ g AlCl}_3$$

$$39.0 \text{ g Cl}_2 \times \frac{1 \text{ mol Cl}_2}{70.906 \text{ g Cl}_2} \times \frac{2 \text{ mol AlCl}_3}{3 \text{ mol Cl}_2} \times \frac{133.341 \text{ g AlCl}_3}{1 \text{ mol AlCl}_3} =$$

Answer: 48.9 g AlCl₃

2. Identify the limiting reagent when 10.8 g water reacts with 4.7 g Na to produce NaOH and H₂.

To solve, set one reactant to the other and compare what is calculated to the actual amount in the reaction to determine the theoretical limiting reactant.

$$10.8 \text{ g H}_2\text{O} = \text{ ? } \text{ g Na}$$

Or

$$4.7 \text{ g Na} = \text{ ? } \text{ g H}_2\text{O}$$



$$10.8 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.015 \text{ g H}_2\text{O}} \times \frac{2 \text{ mol Na}}{2 \text{ mol H}_2\text{O}} \times \frac{22.990 \text{ g Na}}{1 \text{ mol Na}} = 13.8 \text{ g Na}$$

- ∞ This means that 10.8 g will react completely with 13.8 g Na, but only 4.7 g Na was available to this reaction. Therefore, Na is the limiting reactant in this problem.

Limiting Reagents Answer Key Chemistry

Ensheng Dong



Limiting Reagents Answer Key Chemistry:

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TABLE OF CONTENT JEE MAIN ONLINE PAPER 2020 Jan Sep Attempt

Some basic concepts of Chemistry States of Matters Atomic Structure Chemical Bonding Thermodynamics Solutions

Equilibrium Redox Reaction and Electrochemistry Chemical Kinetics and Surface Chemistry Periodicity of Elements Principles and Processes of Metallurgy Hydrogen s Block and p Block Elements d and f block Elements and Coordination Chemistry Environmental Chemistry General Organic Chemistry Hydrocarbons and their Halogen Derivatives Organic Compounds Containing Oxygen Alcohols Ethers Aldehydes ketones Carboxylic Acids and their Derivatives Organic Compounds Containing Nitrogen Amines and Diazonium Salts Polymers and Biomolecules Analytical Chemistry and Chemistry in Daily life Practice Sets for JEE MAIN Practice Sets 1 3 *Introductory Basics Of Chemistry* Dr J. Sai Chandra, 2022-05-07 Our Chemistry Reference Book adheres to the scope and sequence of most general chemistry courses nationwide We strive to make chemistry as a discipline interesting and accessible to students With this objective in mind the content of this Reference Book has been developed and arranged to provide a logical progression from fundamental to more advanced concepts of chemical science Topics are introduced within the context of familiar experiences whenever possible treated with an appropriate rigor to satisfy the intellect of the learner and reinforced in subsequent discussions of related content The organization and pedagogical features were developed and vetted with feedback from chemistry educators dedicated to the project Dr J SAI CHANDRA Mr SANTOSH RAMCHANDRA KSHIRSAGAR Dr SAMBHAJI MAHIPATI KALE Mr SANDIP PANDURANG GONDAKE Mr SAGAR INDRAJEET SHINDE **20 Years Chapterwise Topicwise (2021-2002)**

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aspirants for the forthcoming exam Table of Content Some Basic Concepts of Chemistry Atomic Structure Periodic
Classification and Periodic Properties Chemical Bonding States of Matter Chemical and Ionic Equilibrium Thermodynamics
and Thermochemistry Solid State Solutions and Colligative Properties Electrochemistry Chemical Kinetics Nuclear Chemistry
Surface Chemistry s block Elements p block Elements I p block Elements II Transition and Inner Transition Elements

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Table of Contents Limiting Reagents Answer Key Chemistry

1. Understanding the eBook Limiting Reagents Answer Key Chemistry
 - The Rise of Digital Reading Limiting Reagents Answer Key Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reagents Answer Key Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reagents Answer Key Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reagents Answer Key Chemistry
 - Personalized Recommendations
 - Limiting Reagents Answer Key Chemistry User Reviews and Ratings
 - Limiting Reagents Answer Key Chemistry and Bestseller Lists
5. Accessing Limiting Reagents Answer Key Chemistry Free and Paid eBooks

- Limiting Reagents Answer Key Chemistry Public Domain eBooks
 - Limiting Reagents Answer Key Chemistry eBook Subscription Services
 - Limiting Reagents Answer Key Chemistry Budget-Friendly Options
6. Navigating Limiting Reagents Answer Key Chemistry eBook Formats
- ePub, PDF, MOBI, and More
 - Limiting Reagents Answer Key Chemistry Compatibility with Devices
 - Limiting Reagents Answer Key Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Limiting Reagents Answer Key Chemistry
 - Highlighting and Note-Taking Limiting Reagents Answer Key Chemistry
 - Interactive Elements Limiting Reagents Answer Key Chemistry
8. Staying Engaged with Limiting Reagents Answer Key Chemistry
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reagents Answer Key Chemistry
9. Balancing eBooks and Physical Books Limiting Reagents Answer Key Chemistry
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reagents Answer Key Chemistry
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Limiting Reagents Answer Key Chemistry
- Setting Reading Goals Limiting Reagents Answer Key Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Limiting Reagents Answer Key Chemistry
- Fact-Checking eBook Content of Limiting Reagents Answer Key Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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