



Name: REDACTED

Date: 2/17/2022

Student Exploration: Dehydration Synthesis

Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

Vocabulary: carbohydrate, chemical formula, dehydration synthesis, disaccharide, glucose, hydrolysis, monosaccharide, polysaccharide, valence

Prior Knowledge Questions (Do these **BEFORE** using the Gizmo.)

1. If you exercise on a hot day, you need to worry about dehydration. In this context, what do you think dehydration means?

Dehydration is the lack/loss/or removal of water from something/someone

2. Astronauts and backpackers often bring dehydrated food. What do you think dehydrated food is?

Foods that lack water molecules, if I recall correctly it also preserves the food.

Gizmo Warm-up

What do rice, potatoes, and sugar have in common? They are all foods rich in **carbohydrates**. Carbohydrates are an important energy source for your body. The basic building block of most carbohydrate compounds is the molecule **glucose**. Using the Dehydration Synthesis Gizmo, you will learn about the structure of a glucose molecule and how glucose molecules can be joined together to make larger carbohydrate molecules.

To begin, select the **ORIENT GLUCOSE** tab.

1. Look at the **chemical formula** for glucose. How many carbon (C), hydrogen (H), and oxygen (O) atoms are found in a molecule of glucose? C: 6 H: 12 O: 6

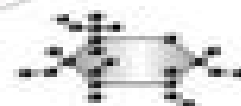
2. Turn on **Show chemical structure**. Each black sphere represents a carbon, hydrogen, or oxygen atom. The lines connecting the spheres represent chemical bonds.

- A. How many black spheres are in the diagram?

24 Black spheres

- B. How does this relate to the number of carbon, hydrogen, and oxygen atoms in the chemical formula for glucose?

The # of spheres is the total number of C, H, and O atoms added up together (12+6+6 = 24)



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Lifelique



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Limiting Reactant and Percent Lifeliqe,2019 This lesson plan covers analyzing chemical reactions in order to determine limiting reactants and excess reactants including calculating the amount of excess reactant calculate the theoretical yield of a reaction and calculating the percent yield of a reaction

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