

Calculations need to be based on the limiting reactant.

- Example 2: What is the mass of sulfur dioxide that can be produced by the reaction of 95.6 g carbon disulfide with 100. g oxygen?



$$\frac{95.6\text{g CS}_2}{76.2\text{g CS}_2} \times \frac{1\text{mol CS}_2}{1\text{mol CS}_2} \times \frac{2\text{mol SO}_2}{1\text{mol CS}_2} \times \frac{64.1\text{g SO}_2}{1\text{mol SO}_2} = 161\text{g SO}_2$$

$$\frac{100.\text{g O}_2}{32.0\text{g O}_2} \times \frac{1\text{mol O}_2}{1\text{mol O}_2} \times \frac{2\text{mol SO}_2}{3\text{mol O}_2} \times \frac{64.1\text{g SO}_2}{1\text{mol SO}_2} = 134\text{g SO}_2$$

134 g of SO_2 can be produced in this reaction.

Limiting Reactant Quick Lab Answers

Ian Guch



Limiting Reactant Quick Lab Answers:

Unveiling the Power of Verbal Art: An Psychological Sojourn through **Limiting Reactant Quick Lab Answers**

In a global inundated with monitors and the cacophony of quick communication, the profound power and emotional resonance of verbal artistry usually diminish in to obscurity, eclipsed by the continuous assault of sound and distractions. Yet, situated within the lyrical pages of **Limiting Reactant Quick Lab Answers**, a captivating function of literary splendor that pulses with fresh feelings, lies an remarkable trip waiting to be embarked upon. Written by a virtuoso wordsmith, that enchanting opus courses readers on an emotional odyssey, delicately exposing the latent potential and profound affect embedded within the complex web of language. Within the heart-wrenching expanse of the evocative evaluation, we will embark upon an introspective exploration of the book is main styles, dissect their fascinating writing style, and immerse ourselves in the indelible effect it leaves upon the depths of readers souls.

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